

Solutions Study Guide

6. What factor increases solubility of a solid?

7. Identify the

solubility

8. Identify the

9. 100 g

10. 100 g

11. 2.5 g

12. Identify the

13. 100 g

14. 100 g

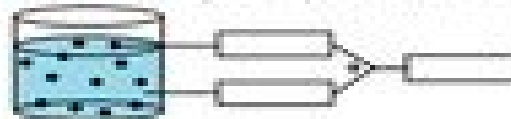
15. 100 g

16. 100 g

Name _____ Period _____ Date _____

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1. Identify the parts of a solution by labeling the following amount. (Solvent, Solute, Solution)



2. Circle which of the following are Solutions & Cross out the ones which are not solutions.

- ☐ Sugar
- ☐ Iron
- ☐ Salt Water
- ☐ Orange Juice
- ☐ Air
- ☐ Brass

3. Water plays an important role in creating solutions. It is known as the _____.

a. Universal Solvent

b. Universal Solvent

c. Why does water make such a good solvent?

d. Which type of solute would be soluble in water? (Remember "like dissolves like")

4. Order a word to match the word to its definition.

Solute	What is dissolved
Solvent	Substance that dissolves another substance in the solution
Solution	Substance that dissolves
Saturated	Substance that cannot be dissolved
Electrolyte	Substance that conducts electricity
Concentration	Amount of solute in a solution
Unsaturated	Substance that dissolves in water and cannot be dissolved
Supersaturated	Substance that dissolves in water and cannot be dissolved
Saturated	Substance that dissolves in water and cannot be dissolved
Unsaturated	Substance that dissolves in water and cannot be dissolved

Solutions Study Guide Chemistry

John William Hill, Dorothy M. Feigl



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