

Chapter 1

The Schrödinger Equation

1.1 (a) F; (b) T; (c) T.

1.2 (a) $E_{\text{photon}} = h\nu = hc/\lambda = (6.626 \times 10^{-34} \text{ J s})(2.998 \times 10^8 \text{ m/s})/(1064 \times 10^{-9} \text{ m}) = 1.867 \times 10^{-19} \text{ J}$.

(b) $E = (5 \times 10^6 \text{ J/s})(2 \times 10^{-8} \text{ s}) = 0.1 \text{ J} = n(1.867 \times 10^{-19} \text{ J})$ and $n = 5 \times 10^{17}$.

1.3 Use of $E_{\text{photon}} = hc/\lambda$ gives

$$E = \frac{(6.022 \times 10^{23})(6.626 \times 10^{-34} \text{ J s})(2.998 \times 10^8 \text{ m/s})}{300 \times 10^{-9} \text{ m}} = 399 \text{ kJ}$$

1.4 (a) $T_{\text{max}} = h\nu - \Phi =$

$$(6.626 \times 10^{-34} \text{ J s})(2.998 \times 10^8 \text{ m/s})/(200 \times 10^{-9} \text{ m}) - (2.75 \text{ eV})(1.602 \times 10^{-19} \text{ J/eV}) = 5.53 \times 10^{-19} \text{ J} = 3.45 \text{ eV}.$$

(b) The minimum photon energy needed to produce the photoelectric effect is $(2.75 \text{ eV})(1.602 \times 10^{-19} \text{ J/eV}) = h\nu = hc/\lambda = (6.626 \times 10^{-34} \text{ J s})(2.998 \times 10^8 \text{ m/s})/\lambda$ and $\lambda = 4.51 \times 10^{-7} \text{ m} = 451 \text{ nm}$.

(c) Since the impure metal has a smaller work function, there will be more energy left over after the electron escapes and the maximum T is larger for impure Na.

1.5 (a) At high frequencies, we have $e^{h\nu/kT} \gg 1$ and the -1 in the denominator of Planck's formula can be neglected to give Wien's formula.

(b) The Taylor series for the exponential function is $e^x = 1 + x + x^2/2! + \dots$. For $x \ll 1$, we can neglect x^2 and higher powers to give $e^x - 1 \approx x$. Taking $x = h\nu/kT$, we have for Planck's formula at low frequencies

$$\frac{d\nu^3}{e^{h\nu/kT} - 1} = \frac{2\pi h\nu^3}{c^2(e^{h\nu/kT} - 1)} \approx \frac{2\pi h\nu^3}{c^2(h\nu/kT)} = \frac{2\pi\nu^2 kT}{c^2}$$

1.6 $\lambda = h/mv = 137h/mc = 137(6.626 \times 10^{-34} \text{ J s})/(9.109 \times 10^{-31} \text{ kg})(2.998 \times 10^8 \text{ m/s}) = 3.32 \times 10^{-10} \text{ m} = 0.332 \text{ nm}$.

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 - The Rise of Digital Reading Solutions Manual Quantum Chemistry Levine
 - Advantages of eBooks Over Traditional Books
2. Identifying Solutions Manual Quantum Chemistry Levine
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solutions Manual Quantum Chemistry Levine
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solutions Manual Quantum Chemistry Levine
 - Personalized Recommendations
 - Solutions Manual Quantum Chemistry Levine User Reviews and Ratings
 - Solutions Manual Quantum Chemistry Levine and Bestseller Lists
5. Accessing Solutions Manual Quantum Chemistry Levine Free and Paid eBooks
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 - Interactive Elements Solutions Manual Quantum Chemistry Levine
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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solutions Manual Quantum Chemistry Levine
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solutions Manual Quantum Chemistry Levine
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Setting Reading Goals Solutions Manual Quantum Chemistry Levine
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Solutions Manual Quantum Chemistry Levine
 - Fact-Checking eBook Content of Solutions Manual Quantum Chemistry Levine
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
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