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# Solutions Vocabulary Review Chemistry Answers

**JS Bruner**



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web limiting reactant theoretical and percentage yields key  $4 KO_2 + 2 H_2O \rightarrow 4 KOH + 3 O_2$  g if a reaction vessel contains 0 mol  $KO_2$  and 0 mol  $H_2O$  what is the limiting reactant

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web limiting reactant reactant present in an amount lower than required by the reaction stoichiometry thus limiting the amount of product generated percent yield measure of the efficiency of a reaction expressed as a percentage of the theoretical yield

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8 3 mole H<sub>2</sub>S and multiply it with 2 mole H<sub>2</sub>O over 2 mole H<sub>2</sub>S from the equation  

$$8 3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8 3 \text{ mole H}_2\text{O}$$
  
 8 3 mole H<sub>2</sub>O  $\times$  18 g H<sub>2</sub>O / 1 mole H<sub>2</sub>O = 149 4 g H<sub>2</sub>O is your theoretical yield so percent yield =  $\frac{137 1 \text{ g H}_2\text{O}}{149 4 \text{ g H}_2\text{O}} \times 100 = 91 77\%$   
 problems 6 8 are similar to problem 5

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web 1 csun edu hcchm001 introchemhandouts html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH<sub>3</sub> can be produced theoretically from the reaction of 5 0 g of N  

$$\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$$
  
 what is the limiting reagent if 8 52 g are actually formed what is the

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web limiting reagents problem one 1 for the balanced equation  

$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$
  
 if the reaction of 21 3 grams of O<sub>2</sub> produces a 47 3 % yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have  

$$21 3 \text{ g O}_2 \times \frac{1 \text{ mole O}_2}{32 \text{ g O}_2} = 0 666 \text{ moles O}_2$$
  

$$0 666 \text{ moles O}_2 \times \frac{3 \text{ moles CO}}{3 \text{ moles O}_2} = 0 666 \text{ moles CO}$$
  

$$0 666 \text{ moles CO} \times 28 \text{ g CO} = 18 6 \text{ g CO}$$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency =  $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$  in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction  

$$\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$$
  
 b determine the theoretical yield of Al<sub>2</sub>O<sub>3</sub> 160 0

g al x 1 mol al x 2 mol al<sub>2</sub>O<sub>3</sub> x 101 964 g al<sub>2</sub>O<sub>3</sub> 302 3 g al<sub>2</sub>O<sub>3</sub> 26 982 g al 4 mol al 1 mol al<sub>2</sub>O<sub>3</sub> c determine the percent yield  
 yield actual x 100 260 0 g x 100 86 01 theoretical 302 3 g

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be  
 weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12  
 c is the limiting reagent 4 33 g of h 2 are left over 13

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 remains when the reaction is complete

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web the percent yield is calculated as follows text percent yield  $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  percent  
 yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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 how much product will be produced from the limiting reactant calculate how much reactant s remains when the reaction is  
 complete

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 ch<sub>3</sub>oh can be produced by the following reaction 2h<sub>2</sub> co ch<sub>3</sub>oh calculate the theoretical yield of ch<sub>3</sub>oh if 68 5 g of co is  
 reacted with 8 6 g of h<sub>2</sub>

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 reaction

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 which reactant is limiting assuming we started with 30 0 grams of ammonium nitrate and 50 0 grams of sodium phosphate  
 what is the

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 80 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2$  x 1 333.8 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  28 g  $\text{CO}$  1 mol  $\text{CO}$

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web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage  
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