

5. Predict which system at equilibrium will contain the maximum amount of product and which system will contain the maximum amount of reactant?



Ans: The magnitude of  $K_c$  can give us some information about the reactant and product concentrations at equilibrium:

- If  $K_c$  is very large,  $\sim 1000$  or more, we will have mostly product species present at equilibrium.
- If  $K_c$  is very small,  $\sim 0.001$  or less, we will have mostly reactant species present at equilibrium.
- If  $K_c$  is between  $0.001$  and  $1000$ , we will have a significant concentration of both reactant and product species present at equilibrium.

By using these guidelines, we can quickly estimate whether a reaction will strongly favour the forward direction to make products very large  $K_c$  – strongly favour the backward direction to make reactants very small  $K_c$  or somewhere in between.

Therefore, system (b) will contain the maximum amount of product.

### SECTION 'D' : NUMERICAL

1. Dinitrogen tetraoxide  $\text{N}_2\text{O}_4$  decomposed into nitrogen dioxide  $\text{NO}_2$  in a reversible reaction. Derive equilibrium constant expression for the reaction of decomposition. Also, interpret the unit of  $K_c$  for balanced chemical reversible reaction.

Ans: Equation:-



Equilibrium Constant Expression:  $K_c = \frac{[\text{NO}_2]^2}{[\text{N}_2\text{O}_4]}$

Unit of  $K_c$  :  $K_c = \frac{[\text{mol}.\text{dm}^{-3}]^2}{[\text{mol}.\text{dm}^{-3}]}$   
 $= \text{mol}.\text{dm}^{-3}$

# Solved Exercises Of Chemistry Of 11th Hsc

**Mike Jess**



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