

Worksheet #3—Stoichiometry

Directions: Solve the following stoichiometry problems. Assume all gases are at STP. Don't forget to write out your balanced chemical equation. Make sure to show ALL work to receive FULL CREDIT!

1. Calculate the number of moles of oxygen required to react completely with 14.8 mol Al in the following equation: $4\text{Al}_{(s)} + 3\text{O}_{2(g)} \rightarrow 2\text{Al}_2\text{O}_{3(s)}$

$$\frac{14.8 \text{ moles Al}}{4 \text{ mole Al}} \times \frac{3 \text{ moles of O}_2}{1 \text{ mole Al}} = \underline{11.1 \text{ mol O}_2}$$

3. If 20.0g grams of zinc react with excess hydrochloric acid, how many grams of zinc chloride (ZnCl_2) are produced?



$$\frac{20 \text{ g Zn}}{65 \text{ g Zn}} \times \frac{1 \text{ mol Zn}}{1 \text{ mole Zn}} \times \frac{1 \text{ mole ZnCl}_2}{1 \text{ mole Zn}} \times \frac{136 \text{ g ZnCl}_2}{1 \text{ mol ZnCl}_2}$$

$$= \underline{41.8 \text{ g ZnCl}_2}$$

5. How many grams of copper are required to replace 4.00 grams of silver nitrate which are dissolved in water?



$$\frac{4.0 \text{ g AgNO}_3}{170 \text{ g AgNO}_3} \times \frac{1 \text{ mol AgNO}_3}{170 \text{ g AgNO}_3} \times \frac{1 \text{ mole Cu}}{2 \text{ mol AgNO}_3} \times \frac{63.5 \text{ g Cu}}{1 \text{ mole Cu}}$$

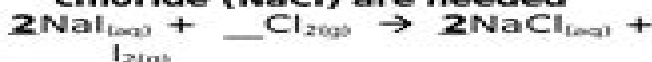
2. Calculate the moles of oxygen produced from the decomposition of 0.61 moles of potassium chlorate (KClO_3).



$$\frac{0.61 \text{ moles KClO}_3}{2 \text{ mole KClO}_3} \times \frac{3 \text{ moles of O}_2}{1 \text{ mole KClO}_3} =$$

$$\underline{0.915 \text{ mol O}_2}$$

4. How many grams of chlorine gas must be reacted with excess sodium iodide (NaI) if 10.0 grams of sodium chloride (NaCl) are needed



$$\frac{10 \text{ g NaCl}}{58.5 \text{ g NaCl}} \times \frac{1 \text{ mol NaCl}}{58.5 \text{ g NaCl}} \times \frac{1 \text{ mol Cl}_2}{2 \text{ mol NaCl}} \times \frac{71 \text{ g Cl}_2}{1 \text{ mol Cl}_2}$$

$$= \underline{6.07 \text{ g Cl}_2}$$

6. How many moles of hydrogen gas are needed to react with 15.1 g chlorine gas to produce hydrogen chloride gas?



$$\frac{15.1 \text{ g Cl}_2}{71 \text{ g Cl}_2} \times \frac{1 \text{ mol Cl}_2}{71 \text{ g Cl}_2} \times \frac{1 \text{ mole H}_2}{1 \text{ mole Cl}_2}$$

$$= \underline{0.21 \text{ moles H}_2}$$

Solving Problems A Chemistry H Answer Key

Mark M. Benjamin



Solving Problems A Chemistry H Answer Key:

A Problem-Solving Approach to Aquatic Chemistry James N. Jensen, 2023-01-05 A Problem Solving Approach to Aquatic Chemistry Enables civil and environmental engineers to understand the theory and application of aquatic equilibrium chemistry The second edition of A Problem Solving Approach to Aquatic Chemistry provides a detailed introduction to aquatic equilibrium chemistry calculation methods for systems at equilibrium applications of aquatic chemistry and chemical kinetics The text directly addresses two required ABET program outcomes in environmental engineering chemistry including stoichiometry equilibrium and kinetics and material and energy balances fate and transport of substances in and between air water and soil phases The book is very student centered with each chapter beginning with an introduction and ending with a summary that reviews the chapter's main points To aid in reader comprehension important terms are defined in context and key ideas are summarized Many thought provoking discussion questions worked examples and end of chapter problems are also included Each part of the text begins with a case study a portion of which is addressed in each subsequent chapter illustrating the principles of that chapter In addition each chapter has an Historical Note exploring connections with the people and cultures connected to topics in the text A Problem Solving Approach to Aquatic Chemistry includes Fundamental concepts such as concentration units thermodynamic basis of equilibrium and manipulating equilibria Solutions of chemical equilibrium problems including setting up the problems and algebraic graphical and computer solution techniques Acid base equilibria including the concepts of acids and bases titrations and alkalinity and acidity Complexation including metals ligands equilibrium calculations with complexes and applications of complexation chemistry Oxidation reduction equilibria including equilibrium calculations graphical approaches and applications Gas liquid and solid liquid equilibrium with expanded coverage of the effects of global climate change Other topics including chemical kinetics of aquatic systems surface chemistry and integrative case studies For advanced senior undergraduates and first year graduate students in environmental engineering courses A Problem Solving Approach to Aquatic Chemistry serves as an invaluable learning resource on the topic with a variety of helpful learning elements included throughout to ensure information retention and the ability to apply covered concepts in practical settings

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edition of Organic Chemistry continues Solomons Fryhle Snyder s tradition of excellence in teaching and preparing students for success in the organic classroom and beyond A central theme of the authors approach to organic chemistry is to emphasize the relationship between structure and reactivity To accomplish this the content is organized in a way that combines the most useful features of a functional group approach with one largely based on reaction mechanisms The authors philosophy is to emphasize mechanisms and their common aspects as often as possible and at the same time use the unifying features of functional groups as the basis for most chapters The structural aspects of the authors approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show students what it does in living systems and the physical world around us Catalog of Copyright Entries. Third Series Library of Congress. Copyright Office,1970 **Chemistry Workbook For Dummies** Peter

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