

Assignment

Date

Period

Solve each equation by factoring.

1) $k^2 + 2k - 48 = 0$

$(k+8)(k-6) = 0$

$k = -8, k = 6$

$\begin{array}{r} -48 \overline{) 12} \\ 2 \cdot 6 \end{array}$

3) $x^2 - 7x + 6 = 0$

$(x-6)(x-1) = 0$

$x = 6, 1$

$\begin{array}{r} 6 \overline{) 7} \\ -6 \overline{) 1} \end{array}$

5) $m^2 - 9m + 18 = 0$

$(m-6)(m-3) = 0$

$m = 6, m = 3$

$\begin{array}{r} 18 \overline{) 9} \\ -6 \overline{) 3} \end{array}$

7) $21x^2 - 38x - 48 = 0$

$\begin{array}{r} -100 \overline{) 38} \\ -38 \end{array}$

$x = \frac{8}{3}, -\frac{6}{7}$

9) $b^2 + 8 = -6b$

$b^2 + 6b + 8 = 0$

$(b+2)(b+4) = 0$

$b = -2, b = -4$

$\begin{array}{r} 8 \overline{) 6} \\ 4 \cdot 2 \end{array}$

Solve each equation by taking square roots.

11) $6 + 81n^2 = 70$

$81n^2 = 64$

$\sqrt{n^2} = \sqrt{\frac{64}{81}}$

$n = \pm \frac{8}{9}$

13) $64x^2 - 4 = 60$

$64x^2 = 64$

$x^2 = 1$

$x = \pm 1$

15) $4p^2 + 2 = 66$

$4p^2 = 64$

$p^2 = 16$

$p = \pm 4$

2) $p^2 + 13p + 42 = 0$

$(p+6)(p+7) = 0$

$p = -6, p = -7$

$\begin{array}{r} 13 \overline{) 42} \\ 1 \cdot 7 \end{array}$

4) $x^2 - 4x + 4 = 0$

$(x-2)(x-2) = 0$

$x = 2, x = 2$

$\begin{array}{r} 4 \overline{) 4} \\ -2 \overline{) 2} \end{array}$

6) $21r^2 - 47r + 20 = 0$

$21r^2 - 35r + 12r + 20 = 0$

$7r(3r-5) + 4(3r-5) = 0$

$7r-4=0 \quad 3r-5=0$

$r = \frac{4}{7}, r = \frac{5}{3}$

$\begin{array}{r} 420 \overline{) 47} \\ -35 \overline{) 12} \end{array}$

8) $n^2 - 8 = -2n$

$n^2 + 2n - 8 = 0$

$(n+4)(n-2) = 0$

$n = -4, n = 2$

$\begin{array}{r} -8 \overline{) 2} \\ -4 \overline{) 2} \end{array}$

10) $v^2 = 35 - 2v$

$v^2 + 2v - 35 = 0$

$(v+7)(v-5) = 0$

$v = -7, v = 5$

$\begin{array}{r} 35 \overline{) 2} \\ 7 \overline{) 5} \end{array}$

12) $3 - 8a^2 = -125$

$-8a^2 = -128$

$a^2 = 16$

$a = \pm 4$

14) $7 - 2k^2 = -65$

$-2k^2 = -72$

$k^2 = 36$

$k = \pm 6$

Solving Quadratic Equations Test Answers

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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g}) \quad \Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \quad \Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g}) \quad \Delta h = 283 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g})$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l}) \quad \Delta h = 1411 \text{ kJ/mole}$
 $\text{C}_2\text{H}_6(\text{g}) + 7/2 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l}) \quad \Delta h = 1560 \text{ kJ/mole}$
 $\text{H}_2(\text{g}) + 1/2 \text{ O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta h = 285 \text{ kJ/mole}$

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values to determine the heat of reaction for the reaction C_2H_4

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web thus for reaction 9 2 in which 1 mole of H_2O and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = -55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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