

Esempio: calcolare quante molecole sono contenute in un bicchier d'acqua avente la capacità di 0,135 litri.

L'acqua presenta una densità di 1 kg dm^{-3} ($1 \text{ dm}^3 = 1 \text{ l}$) e 0,135 l pesano pertanto 135g che corrispondono a

$$n_{H_2O} = \frac{W_{H_2O}}{PM_{H_2O}} = \frac{135g}{18g / mol} = 7,5 mol$$

il numero di molecole presenti in 7,5 moli sarà dunque pari a $7,5 \text{ mol} \times 6,022 \cdot 10^{23} \text{ mol}^{-1} = 4,52 \cdot 10^{24}$

Una conseguenza del principio di Avogadro è che un medesimo numero di moli di una qualsiasi sostanza gassosa devono occupare sempre il medesimo volume (a P e T costanti). Infatti se *"volumi uguali di gas diversi nelle stesse condizioni di T e P contengono lo stesso numero di particelle"*, allora deve anche essere vero che *"gas che contengano lo stesso numero di particelle devono occupare lo stesso volume"*. Se ne deduce pertanto che 1 mole di un qualsiasi gas, contenendo sempre lo stesso numero di particelle (il numero di Avogadro) deve occupare sempre il medesimo volume ed in particolare, a 0°C e alla pressione di 1 atm occupa un volume pari a **22,414 l**, detto **volume molare standard**.

1.9 Calcolo delle quantità che reagiscono

Come abbiamo già avuto modo di dire il concetto di mole è essenziale per semplificare i calcoli quantitativi o stechiometrici. La **stechiometria** (dal greco *στοιχειον* = elemento, sostanza fondamentale) è quella parte della chimica che si occupa degli aspetti quantitativi delle reazioni ed in particolare dei **rapporti numerici e ponderali** che intercorrono tra le specie chimiche che reagiscono.

Per poter procedere con i calcoli stechiometrici è necessario che una equazione sia bilanciata. In un'equazione bilanciata sono già definiti i rapporti numerici tra specie chimiche. Quando scriviamo l'equazione bilanciata



individuiamo ad esempio il rapporto numerico di reazione tra Idrogeno e Azoto che risulta essere pari a 3:1. Il rapporto tra Idrogeno ed ammoniaca è invece di 3:2 e così via.

Ma i coefficienti stechiometrici non rappresentano solo il numero di molecole, ma anche il numero di moli che reagiscono. Per rendercene conto, dopo aver ricordato che un'eguaglianza rimane tale se moltiplichiamo entrambi i suoi membri per uno stesso numero, moltiplichiamo entrambi i membri per il numero di Avogadro (N_A).

$$N_A \cdot (3\text{H}_2 + \text{N}_2) = N_A \cdot (2\text{NH}_3)$$

Applichiamo la proprietà distributiva:

$$3 N_A \cdot (\text{H}_2) + 1 N_A \cdot (\text{N}_2) = 2 N_A \cdot (\text{NH}_3)$$

ma per la definizione di mole: un numero di Avogadro di molecole di Azoto costituiscono una mole di Azoto, 3 numeri di Avogadro di molecole di Idrogeno costituiscono 3 moli di Idrogeno e 2 numeri di Avogadro di molecole di Ammoniaca costituiscono 2 moli di Ammoniaca.

Se ne deduce quindi che i rapporti numerici precedentemente individuati tra i coefficienti stechiometrici non sono solo rapporti molecolari, ma anche rapporti molari. In altre parole per ogni 3 moli di Idrogeno reagisce 1 mole di Azoto per dare 2 moli di Ammoniaca.

Sommario Sommario Teoria Atomica E Leggi Quantitative 1

Michael Brown



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Table of Contents Sommario Sommario Teoria Atomica E Leggi Quantitative 1

1. Understanding the eBook Sommario Sommario Teoria Atomica E Leggi Quantitative 1
 - The Rise of Digital Reading Sommario Sommario Teoria Atomica E Leggi Quantitative 1
 - Advantages of eBooks Over Traditional Books
2. Identifying Sommario Sommario Teoria Atomica E Leggi Quantitative 1
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Sommario Sommario Teoria Atomica E Leggi Quantitative 1
 - User-Friendly Interface
4. Exploring eBook Recommendations from Sommario Sommario Teoria Atomica E Leggi Quantitative 1
 - Personalized Recommendations
 - Sommario Sommario Teoria Atomica E Leggi Quantitative 1 User Reviews and Ratings
 - Sommario Sommario Teoria Atomica E Leggi Quantitative 1 and Bestseller Lists
5. Accessing Sommario Sommario Teoria Atomica E Leggi Quantitative 1 Free and Paid eBooks
 - Sommario Sommario Teoria Atomica E Leggi Quantitative 1 Public Domain eBooks
 - Sommario Sommario Teoria Atomica E Leggi Quantitative 1 eBook Subscription Services
 - Sommario Sommario Teoria Atomica E Leggi Quantitative 1 Budget-Friendly Options
6. Navigating Sommario Sommario Teoria Atomica E Leggi Quantitative 1 eBook Formats
 - ePub, PDF, MOBI, and More
 - Sommario Sommario Teoria Atomica E Leggi Quantitative 1 Compatibility with Devices
 - Sommario Sommario Teoria Atomica E Leggi Quantitative 1 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Sommario Sommario Teoria Atomica E Leggi Quantitative 1
 - Highlighting and Note-Taking Sommario Sommario Teoria Atomica E Leggi Quantitative 1
 - Interactive Elements Sommario Sommario Teoria Atomica E Leggi Quantitative 1

8. Staying Engaged with Sommario Sommario Teoria Atomica E Leggi Quantitative 1
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Sommario Sommario Teoria Atomica E Leggi Quantitative 1
9. Balancing eBooks and Physical Books Sommario Sommario Teoria Atomica E Leggi Quantitative 1
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Sommario Sommario Teoria Atomica E Leggi Quantitative 1
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Sommario Sommario Teoria Atomica E Leggi Quantitative 1
 - Setting Reading Goals Sommario Sommario Teoria Atomica E Leggi Quantitative 1
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Sommario Sommario Teoria Atomica E Leggi Quantitative 1
 - Fact-Checking eBook Content of Sommario Sommario Teoria Atomica E Leggi Quantitative 1
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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