

# Chemistry

## Unit 2

### Practice Exam 1 Answers

#### Section A—Multiple-choice questions

##### Question 1

- B** The high electronegativity difference between hydrogen and oxygen means that the O—H bond in water is highly polar, enabling hydrogen bonds to be formed between nearby water molecules.

Sulfur is less electronegative than oxygen. The polarity of the H—S bond in  $\text{H}_2\text{S}$  is much less than the polarity of the O—H bond in water. The attraction between  $\text{H}_2\text{S}$  molecules is described as a dipole-dipole interaction.

The bond between hydrogen and Te is non-polar. The only interactions between molecules of  $\text{H}_2\text{Te}$  are dispersion forces brought about through the formation of instantaneous dipoles.

##### Question 2

- D** The formation of a solution involves breaking bonds between solute particles, breaking intermolecular bonds between solvent molecules, and the formation of bonds between solvent molecules and solute particles.

##### Question 3

- C**  $n(\text{Na}_2\text{CO}_3) = cV = 0.150 \times \frac{300}{1000} = 0.0450 \text{ mol}$   
 $m(\text{Na}_2\text{CO}_3) = mM = 0.0450 \times 106.0 = 4.77 \text{ g}$

##### Question 4

- C** Water is a poor conductor of electricity due to the small number of hydronium ions and hydroxide ions that are produced by the self-ionisation reaction of water.

##### Question 5

- A** Nitric acid,  $\text{HNO}_3$ , is a strong acid because it is completely ionised in water. Ethanoic acid,  $\text{CH}_3\text{COOH}$ , is a weak acid as it is partly ionised in water. A 0.1 M solution is considered to be dilute and 6 M concentrated.

##### Question 6

- B** When the acid  $\text{H}_2\text{PO}_4^-$  donates a proton, it forms the conjugate base  $\text{HPO}_4^{2-}$ .

##### Question 7

- A** pH is defined as  $\text{pH} = -\log_{10}[\text{H}_3\text{O}^+]$ . Since HCl is a strong acid and completely ionised in solution, it can be assumed that in each beaker  $[\text{HCl}] = [\text{H}_3\text{O}^+]$ . pH is only dependent on the  $[\text{H}_3\text{O}^+]$ , not the volume of solution. Because the definition involves the negative logarithm, the pH increases as the  $[\text{H}_3\text{O}^+]$  decreases. The solution that has the lowest  $[\text{H}_3\text{O}^+]$  will have the highest pH.

# Pearson Successnet Chemistry Answers

**Susan McMurry**



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