# Science

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(Chapter - 2) (Acids, Bases and Salts) (Exercises)

#### (Class 10)

$$3H_2SO_{4(aq)} + 2Al_{(s)} \longrightarrow Al_2(SO_4)_{3(aq)} + 3H_{2(g)}$$
  
Sulphuric acid Zinc Aluminium sulphate Hydrogen

$$6HCl_{(aq)}$$
 +  $2Fe_{(s)}$   $\rightarrow$   $2FeCl_{3(aq)}$  +  $3H_{2(g)}$   
Hydrochloric acid Magnesium Magnesium chloride Hydrogen

#### Question 6:

Compounds such as alcohols and glucose also contain hydrogen but are not categorised as acids. Describe an Activity to prove it.

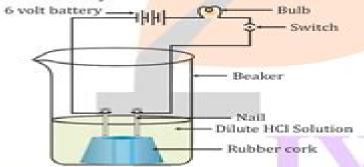
#### LAnswer 6:

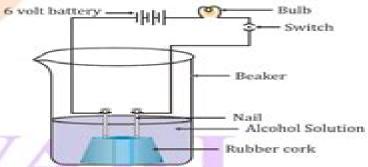
Two nails are fitted on a cork and are kept it in a 100 mL beaker. The nails are then connected to the two terminals of a 6-volt battery through a bulb and a switch. Some dilute HCl is poured in the beaker and the current is switched on. The same experiment is then performed with glucose solution and alcohol solution. Observations

It will be observed that the bulb g lows in the HCl solution and does not g low in the glucose solution.

#### Result

HCl dissociates into H+ and Cl- ions. These ions conduct electricity in the solution resulting in the glowing of the bulb. On the other hand, the glucose solution does not dissociate into ions. Therefore, it does not conduct electricity.





#### Conclusion

From this activity, it can be concluded that all acids contain hydrogen but not all compounds containing hydrogen are acids. That is why, though alcohols and glucose contain hydrogen, they are not categorised as acids.

#### Question 7:

Why does distilled water not conduct electricity, whereas rain water does?

#### Answer 7:

Distilled water is a pure form of water not containing any ionic species. Therefore, it does not conduct electricity. Rain water, being an impure form of water, contains many ionic species (dissolved in air) such as acids and therefore it conducts electricity.

# **Science 10 Chapter 2 Test**

Jane R. Burstein, Carolyn C. Wheater

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