

84 grams Hydrogen (H₂) and 1 mole of oxygen (O₂)? Review! Reaction Stoichiometry
 $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$

72. Write the dissociation equation for Li₂SO₄. If no reaction, state why.

73. Write the dissociation equation for AgCl. If no reaction, state why.



74. Write the dissociation equation for KOH. If no reaction, state why.

75. 3 How many ions are formed in the dissociation of MgF₂ in water?

76. What is the molarity of a solution that contains 0.123 moles KCl in 4.54L solution? see notes

$$M = \frac{\text{mol}}{L} \quad \frac{0.123 \text{ mol}}{4.54 \text{ L}} = \boxed{0.027 \text{ M}}$$

77. How many moles of HCl are present in 0.500 L of a 0.1 M HCl solution? see notes

$$M = \frac{\text{mol}}{L} \quad 0.15 = \frac{\text{mol}}{0.50 \text{ L}} \quad \boxed{0.14 \text{ mol HCl}}$$

78. What volume of concentrated sulfuric acid, 6.5M, is required to prepare 2.0L of 0.25M solution by dilution with water? see notes

$$\begin{aligned} V_1 &= ? & V_2 &= 2.00 \text{ L} \\ M_1 &= 6.5 \text{ M} & M_2 &= 0.25 \text{ M} \end{aligned} \quad \begin{aligned} 6.5V &= (2)(0.25) \\ 6.5V &= 0.5 \\ \frac{6.5V}{6.5} &= \frac{0.5}{6.5} \end{aligned} \quad \boxed{V = 0.077 \text{ L}} \\ &&& \text{or } 77 \text{ mL}$$

79. If 1.4 L of a 2.2 M solution is diluted to 1.1 M solution, what is the new volume? see notes

$$\begin{aligned} V_1 &= 1.4 \text{ L} & V_2 &= ? \\ M_1 &= 2.2 \text{ M} & M_2 &= 1.5 \text{ M} \end{aligned} \quad \begin{aligned} (1.4)(2.2) &= \frac{1.5V}{1.5} \\ \frac{3.08}{1.5} &= \frac{1.5V}{1.5} \end{aligned} \quad \boxed{V = 2.1 \text{ L}}$$

80. How many grams of a Na₂C₂H₃O₂ are needed to prepare 100 mL of a 0.5 M solution? see notes

$$\begin{aligned} M &= \frac{\text{mol}}{L} \\ 0.5 \text{ M} &= \frac{\text{mol}}{0.1 \text{ L}} \end{aligned} \quad \frac{0.33 \text{ mol NaC}_2\text{H}_3\text{O}_2}{1 \text{ mol NaC}_2\text{H}_3\text{O}_2} \quad \frac{32.04 \text{ g NaC}_2\text{H}_3\text{O}_2}{1 \text{ mol NaC}_2\text{H}_3\text{O}_2} \quad \boxed{27 \text{ g}}$$

81. At what temperature will 0.234 mol of CO₂ exert a pressure of 4.56 atm in a volume of 0.234 L? see notesReview! Unit 9: Gas Laws

$$PV = nRT$$

$$(4.56)(0.234) = (0.234)(0.0821)T$$

$$\frac{1.06704}{0.0101} = \frac{0.0191014}{0.0101} T$$

$$\boxed{T = 105.6 \text{ K}}$$

83. During a lab experiment, students produced 15.6 grams Gold (Au), however, 26.3 grams (100% Au) could have theoretically been produced. What is the percent yield of this lab experiment? see notes

Review! Reaction Stoichiometry: Theoretical yield
 $\% \text{ yield} = \frac{\text{actual}}{\text{theoretical}} \times 100$

$$\% \text{ yield} = \frac{15.6 \text{ g}}{26.3 \text{ g}} = \boxed{59.3\% \text{ yield}}$$

84. Which of the following is/are precipitation reaction(s). Circle the precipitate if applicable.



$\rightarrow 2\text{NaOH(aq)} + \text{Na}_2\text{CO}_3\text{(aq)} \rightarrow 2\text{NaOH(aq)} + \text{Na}_2\text{CO}_3\text{(aq)}$
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 all products soluble

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