

Read This!

Electronegativity is a measure of the ability of an atom's nucleus to attract electrons from a different atom within a covalent bond. A higher electronegativity value correlates to a stronger pull on the electrons in a bond. This value is only theoretical. It cannot be directly measured in the lab.

12. Using the definition stated in the *Read This!* box above, select the best visual representation for electronegativity. Explain your reasoning.



13. Locate the electronegativity values in Model 1.

- a. What is the trend in electronegativity going down a group in Model 1?

Electronegativity decreases as you go down a group (vertical column).

- b. Explain the existence of the trend described in part a in terms of atomic structure and Coulombic attraction.

The ability to attract electrons is based upon attractive force. The greater the distance between the nucleus and valence electrons, the less the attractive force. Less force means a lower electronegativity.

- c. What is the trend in electronegativity going across a period in Model 1?

Electronegativity increases from left to right across a period (horizontal row)

- d. Explain the existence of the trend described in part c in terms of atomic structure and Coulombic attraction.

The greater the number protons in an atom, in a particular period, the greater the attractive force between the nucleus and valence electrons. The greater attractive force means a higher electronegativity.

14. The two diagrams below can summarize each of the three trends discussed in this activity. Write "atomic radius," "ionization energy," and "electronegativity" under the appropriate diagram.



Pogil Activities For High School Chemistry Periodic Trends Answers

Doctor Chemistry



Pogil Activities For High School Chemistry Periodic Trends Answers:

Atoms, Molecules & Elements: What Are Compounds? Gr. 5-8 George Graybill, 2015-10-01 This is the chapter slice What Are Compounds from the full lesson plan Atoms Molecules Elements Young scientists will be thrilled to explore the invisible world of atoms molecules and elements Our resource provides ready to use information and activities for remedial students using simplified language and vocabulary Students will label each part of the atom learn what compounds are and explore the patterns in the periodic table of elements to find calcium Ca chlorine Cl and helium He through hands on activities These and more science concepts are presented in a way that makes them more accessible to students and easier to understand Written to grade and using simplified language and vocabulary and comprised of reading passages student activities crossword word search comprehension quiz and color mini posters our resource can be used effectively for test prep and your whole class All of our content is aligned to your State Standards and are written to Bloom's Taxonomy and STEM initiatives *High School Chemistry Students' Learning of the Elements, Structure, and Periodicity of the Periodic Table*, 2003 **Atoms and the Periodic Table (Inorganic Chemistry Fast Facts)** E Staff, Learn and review on the go Use Quick Review Chemistry Study Notes to help you learn or brush up on the subject quickly You can use the review notes as a reference to understand the subject better and improve your grades Easy to remember facts to help you perform better Important facts that you need to know about the Periodic Table and atomic trends Perfect study notes for all high school health sciences premed medical and nursing students **Chemistry Lesson Plans, Study Guides, and Lecture Notes** Doctor Chemistry, 2017-08-13 The Ultimate Guide to Learning or Teaching Chemistry This book contains the real lecture notes and slide of a highly effective high school and college Chemistry teacher Teachers Never plan another lesson again Students Ace your upcoming exam This series covers all of the topics of High School Chemistry and General Chemistry including Accuracy and Significant Figures Mixtures Metric System Bonding Atomic Theory Periodic Table VSEPR Ionic and Covalent Bonding Geometric Bonding The Mole and Molar Mass Equation Balancing Thermodynamics Stoichiometry States of Matter Gas Laws and Calculations Reaction Calculations Acids and Bases Limiting Reagents Redox and Electro Chemistry Organic Chemistry Basics **Simply Science**, 1997 **Integrated Physics and Chemistry, Chapter 1, Activities** Paradigm Accelerated Curriculum, 2005-01-01 Key topics Periodic Table of the Elements metals nonmetals compounds formulas atomic weights heat measuring temperatures Robert Boyle Democritus Lavoisier Proust Dalton Rumford IPC consists of twelve chapters of text and twelve companion student activity books This course introduces students to the people places and principles of physics and chemistry It is written by internationally respected scientist author John Hudson Tiner who applies the vignette approach which effectively draws readers into the text and holds attention The author and editors have deliberately avoided complex mathematical equations in order to entice students into high school level science Focus is on the people who contributed to development of the Periodic Table of the Elements Students learn to read

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