

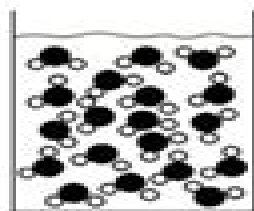
Intermolecular Forces

(What Determines the Boiling Point?)

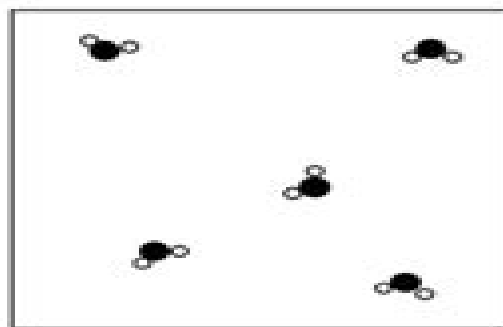
Model 1: Intermolecular Forces in Liquids and Gases.

Molecules attract each other, and the force of attraction increases rapidly as the intermolecular distance decreases. In a liquid, the molecules are very close to one another and are constantly moving and colliding. When a liquid evaporates, molecules in the liquid must overcome these intermolecular attractive forces and break free into the gas phase, where on average molecules are very far apart. For example, when water evaporates, rapidly moving H_2O molecules at the surface of the liquid pull away from neighboring H_2O molecules and enter the gas phase, as shown in Figure 1.

Figure 1. H_2O molecules in the liquid and gas phases.



Attractive forces between water molecules are stronger in the liquid than in the gas because the molecules are very close to one another.



On average, gaseous H_2O molecules are much further apart than in the liquid.

Critical Thinking Questions

1. When water evaporates, are any bonds between H atoms and O atoms within a molecule broken?
2. On average, are the intermolecular forces stronger in $\text{H}_2\text{O}(\ell)$ or in $\text{H}_2\text{O}(\text{g})$? Explain.

Pogil Chemactivity 27 Intermolecular Forces Activity

Kathleen Armour



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