Name	Role(s)	Date

## Naming Ionic Compounds

What are the structural units that make up ionic compounds and how are they named?

## Why?

When working in chemistry, it is often convenient to write a chemical in symbols. For example we might write down the substance table salt as NaCl. In talking about chemistry however, it is a bit tacky to say "en-ay see-ell" when we want to refer to a substance. Also, in formal writing we should use the name of the compound rather than its symbols. Therefore we need to learn how to say the proper names of ionic substances.

## Model 1 - Ion Charges for Selected Elements

Li*	Be3-							N3-	O2 -	Pi-
Nat	Mg2*	Transition elements				AID+		Po-	\$3	CI
K-	Ca <sup>3+</sup>	Fe <sup>3+</sup>	Ni <sup>3+</sup> Ni <sup>3+</sup>	Cu* Cu3*	Zn2-					Brt
Rb.	Rb* Sr2+			Ag1-			Sn <sup>3</sup> * Sn <sup>4</sup> *			I.
	Ba <sup>3</sup>				Hg <sub>2</sub> <sup>3</sup> ·		Pb <sup>3+</sup> Pb <sup>4+</sup>			

1. Based on the information in Model 1:

a. Identify three elements that form only one cation.

b. Identify three elements that form only one anion.

c. Identify three elements that form more than one cation.

d. In what region of the periodic table are these "multiple ion" elements usually located?

- Consider the ions of potassium (K) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(s) of potassium (K) and sulfur (S).
   Keep in mind that the sum of the charges in an ionic compound must equal zero.
- Consider the ions of iron (Fe) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(s) of iron (Fe) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.

# **Pogil Chemistry Batteries Answer Key**

**Léo Pomar** 

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Batteries of Planté and Faure John Hall GLADSTONE (and TRIBE (Alfred)),1883

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## **Table of Contents Pogil Chemistry Batteries Answer Key**

- 1. Understanding the eBook Pogil Chemistry Batteries Answer Key
  - The Rise of Digital Reading Pogil Chemistry Batteries Answer Key
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Pogil Chemistry Batteries Answer Key
  - Exploring Different Genres
  - o Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Pogil Chemistry Batteries Answer Key
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Pogil Chemistry Batteries Answer Key
  - Personalized Recommendations
  - Pogil Chemistry Batteries Answer Key User Reviews and Ratings
  - Pogil Chemistry Batteries Answer Key and Bestseller Lists
- 5. Accessing Pogil Chemistry Batteries Answer Key Free and Paid eBooks
  - Pogil Chemistry Batteries Answer Key Public Domain eBooks
  - o Pogil Chemistry Batteries Answer Key eBook Subscription Services
  - Pogil Chemistry Batteries Answer Key Budget-Friendly Options

- 6. Navigating Pogil Chemistry Batteries Answer Key eBook Formats
  - ∘ ePub, PDF, MOBI, and More
  - Pogil Chemistry Batteries Answer Key Compatibility with Devices
  - Pogil Chemistry Batteries Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Pogil Chemistry Batteries Answer Key
  - Highlighting and Note-Taking Pogil Chemistry Batteries Answer Key
  - Interactive Elements Pogil Chemistry Batteries Answer Key
- 8. Staying Engaged with Pogil Chemistry Batteries Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Pogil Chemistry Batteries Answer Key
- 9. Balancing eBooks and Physical Books Pogil Chemistry Batteries Answer Key
  - Benefits of a Digital Library
  - o Creating a Diverse Reading Collection Pogil Chemistry Batteries Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Pogil Chemistry Batteries Answer Key
  - Setting Reading Goals Pogil Chemistry Batteries Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pogil Chemistry Batteries Answer Key
  - Fact-Checking eBook Content of Pogil Chemistry Batteries Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements

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