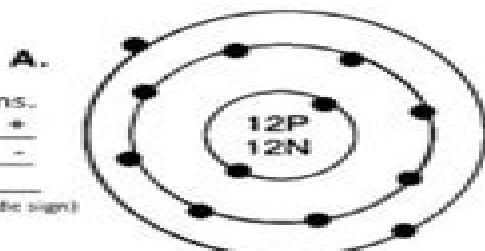
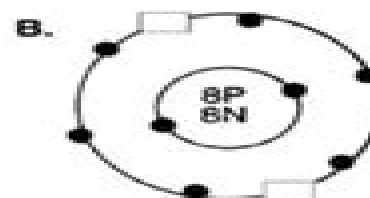


Types of Bonds

Look at the Bohr Models below, move the electrons from one model to the other as they would if these two atoms formed a bond.



After moving the electrons.
How many Protons ____ +
How many Electrons ____ -
What is the difference ____
(include sign)



After moving the electrons.
How many Protons ____ +
How many Electrons ____ -
What is the difference ____

- If atom A loses electrons to atom B,
 - how many electrons will atom A lose? _____
 - how many electrons will atom B gain? _____
 - what will be the oxidation number of atom A? _____ (include sign)
 - what will be the oxidation number of atom B? _____ (include sign)
 - what will be the total charge of the compound formed? Number format.

f. what type of bond will form?

- Explain why an element's oxidation number is related to the group on the periodic table to which it belongs.

For Families 1, 2, 3, and 4 the oxidation number is _____

For Families 5, 6, 7, and 8 the oxidation number is _____

Directions: Complete the table comparing ionic compounds and covalent compounds.

Characteristic	Ionic compounds	Covalent compounds
3. How the compound is formed	electrons are shared electrons are transferred	electrons are shared electrons are transferred
4. Smallest particle	proton, neutron, electron	proton, neutron, electron
5. Usual state at room temperature	Solid, Liquid, Gas	Solid, Liquid, Gas

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AN Whitehead



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Table of Contents Where Are Ionic Bonds Worksheet 233 Physical Science Answers

1. Understanding the eBook Where Are Ionic Bonds Worksheet 233 Physical Science Answers
 - The Rise of Digital Reading Where Are Ionic Bonds Worksheet 233 Physical Science Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Where Are Ionic Bonds Worksheet 233 Physical Science Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Where Are Ionic Bonds Worksheet 233 Physical Science Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Where Are Ionic Bonds Worksheet 233 Physical Science Answers

- Personalized Recommendations
 - Where Are Ionic Bonds Worksheet 233 Physical Science Answers User Reviews and Ratings
 - Where Are Ionic Bonds Worksheet 233 Physical Science Answers and Bestseller Lists
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 - Where Are Ionic Bonds Worksheet 233 Physical Science Answers eBook Subscription Services
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- ePub, PDF, MOBI, and More
 - Where Are Ionic Bonds Worksheet 233 Physical Science Answers Compatibility with Devices
 - Where Are Ionic Bonds Worksheet 233 Physical Science Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Where Are Ionic Bonds Worksheet 233 Physical Science Answers
 - Highlighting and Note-Taking Where Are Ionic Bonds Worksheet 233 Physical Science Answers
 - Interactive Elements Where Are Ionic Bonds Worksheet 233 Physical Science Answers
8. Staying Engaged with Where Are Ionic Bonds Worksheet 233 Physical Science Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Where Are Ionic Bonds Worksheet 233 Physical Science Answers
9. Balancing eBooks and Physical Books Where Are Ionic Bonds Worksheet 233 Physical Science Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Where Are Ionic Bonds Worksheet 233 Physical Science Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Where Are Ionic Bonds Worksheet 233 Physical Science Answers
- Setting Reading Goals Where Are Ionic Bonds Worksheet 233 Physical Science Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Where Are Ionic Bonds Worksheet 233 Physical Science Answers

- Fact-Checking eBook Content of Where Are Ionic Bonds Worksheet 233 Physical Science Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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