Name:______ Date: _____ Period:_____

POGIL- Stoichiometry

How do chemists use balanced chemical equations?

What happened to Avogadro when he got bit by 6.02 x 10 23 mosquitoes? He got Mol-aria

Why?

Chemists use balanced chemical equations as a basis to calculate how much reactant is needed or product is formed in a reaction. This is called <u>Stoichiometry</u>- (stoi-key-ah-meh-tree) Another way of looking at it is using the <u>mole ratio</u> from the balanced equation and information about one compound in the reaction to determine information about another compound in the equation. A <u>mole ratio</u> is a conversion factor <u>derived from the coefficients of a balanced chemical equation</u> interpreted in terms of moles. In chemical calculations, mole ratios are used to convert between moles one thing and moles of another.

Purpose: In this activity we will address the question: How do I convert between different chemical species in a given reaction?

MODEL 1: Mole-to-Mole Stoichiometry (1-step)

Example: mole-to-mole

Question: Given the following equation:

 $1 N_2(g) + 3 H_2(g) \rightarrow 2 NH_3(g)$

How many moles of H2 are needed to produce 34.8 moles of NH1?

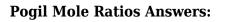
Answer:

Key Questions:

- What was the given in the example above?
- 2. Where did the $\frac{3 \text{ mot} H_2}{2 \text{ mot} N H_3}$, come from?
- Double check to see if 52.2 moles of H2 is the correct answer. SHOW the math below.

Pogil Mole Ratios Answers

RJ Shavelson



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