

Calculating pH

How will it be used for the concentration differences here?

Why?

As buffers and other substances control pH in organisms, it is important to understand the ability of buffers to resist change. This property is called **buffering capacity** and is usually expressed as millimoles per liter of buffering agent or as millimoles per liter of acid or base that can be added to 1 L of solution before the pH of the solution changes by one unit. Buffering capacity is not a property of buffers only; it is a property of any solution.

Table 1 – Ion Concentrations for Acids and Bases

| Number | Substance | Acidity, Alkalinity or Neutrality | $[H^+]$ (M) | $[OH^-]$ (M) | $[H^+] > [OH^-]$ or $[OH^-] > [H^+]$ |
|--------|-------------------------------------|-----------------------------------|-------------------------|-------------------------|--------------------------------------|
| 1 | 0.1 M HCl (strong acid) | Acidic | 1.0×10^{-1} M | 1.0×10^{-13} M | Acidic |
| 2 | 0.001 M HCl (strong acid) | Acidic | 1.0×10^{-3} M | 1.0×10^{-11} M | Acidic |
| 3 | 0.00001 M HCl (strong acid) | Acidic | 1.0×10^{-5} M | 1.0×10^{-9} M | Acidic |
| 4 | 0.0000001 M HCl (strong acid) | Acidic | 1.0×10^{-7} M | 1.0×10^{-7} M | Acidic |
| 5 | 0.00000001 M HCl (strong acid) | Acidic | 1.0×10^{-8} M | 1.0×10^{-6} M | Acidic |
| 6 | 0.000000001 M HCl (strong acid) | Acidic | 1.0×10^{-9} M | 1.0×10^{-5} M | Acidic |
| 7 | 0.0000000001 M HCl (strong acid) | Acidic | 1.0×10^{-10} M | 1.0×10^{-4} M | Acidic |
| 8 | 0.00000000001 M HCl (strong acid) | Acidic | 1.0×10^{-11} M | 1.0×10^{-3} M | Acidic |
| 9 | 0.000000000001 M HCl (strong acid) | Acidic | 1.0×10^{-12} M | 1.0×10^{-2} M | Acidic |
| 10 | 0.0000000000001 M HCl (strong acid) | Acidic | 1.0×10^{-13} M | 1.0×10^{-1} M | Acidic |
| 11 | 0.1 M NaOH (strong base) | Basic | 1.0×10^{-13} M | 1.0×10^{-1} M | Basic |
| 12 | 0.001 M NaOH (strong base) | Basic | 1.0×10^{-11} M | 1.0×10^{-3} M | Basic |
| 13 | 0.00001 M NaOH (strong base) | Basic | 1.0×10^{-9} M | 1.0×10^{-5} M | Basic |
| 14 | 0.0000001 M NaOH (strong base) | Basic | 1.0×10^{-7} M | 1.0×10^{-7} M | Basic |
| 15 | 0.000000001 M NaOH (strong base) | Basic | 1.0×10^{-5} M | 1.0×10^{-9} M | Basic |
| 16 | 0.0000000001 M NaOH (strong base) | Basic | 1.0×10^{-3} M | 1.0×10^{-11} M | Basic |
| 17 | 0.00000000001 M NaOH (strong base) | Basic | 1.0×10^{-1} M | 1.0×10^{-13} M | Basic |

1. Which one of the listed $[H^+]$ is listed in the table?

2. Which one of the listed $[OH^-]$ is listed in the table?

3. For Solution 1, which one has a higher concentration, hydrochloric acid or hydroxide ions?

4. Which one?

5. Which one has the concentration of hydrochloric acid ions listed in the table? 1 or 2? 3 or 4?

6. Which one has the concentration of hydroxide ions listed in the table? 1 or 2? 3 or 4?

7. Which one has the concentration of hydrochloric acid ions listed in the table? 1 or 2? 3 or 4?

8. Which one has the concentration of hydroxide ions listed in the table? 1 or 2? 3 or 4?

Pogil Activities Calculating Ph Answers

**Juliette Lantz, Renée Cole, The POGIL
Project**



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