

Nuclear Decay

Types of Radiation

1. Complete the table below to compare the properties of alpha, beta, and gamma radiation.

Name	Alpha	Beta	Gamma
Greek letter	α	β	γ
Symbol	${}^4_2\text{He}$	${}^0_{-1}\text{e}$	γ
Composition	2 protons, 2 neutrons	an electron	radiation (light)
Charge	0	-1	0
Stopped by	paper, skin, clothing	aluminum	lead

Nuclear Decay Reactions

In problems 2 – 5, fill in the missing particle for each nuclear decay reaction. Then, match each equation to the appropriate type of nuclear decay.



A alpha emission



Fig. 1. *Isotopic composition*



C electron capture

In problems 6 and 7, write a balanced equation for each nuclear decay reaction.

6. Decay of radium-223 by alpha (α) emission



7. Decay of iodine-126 by electron (β) emission



3. How are the atomic number and mass number of a nuclide affected by Beta Decay?

The mass number does not change but the atomic number increases by 1

9. How are the atomic number and mass number of a nuclide affected by the emission of gamma radiation?

The atomic number and mass number do not change

10. Which has a slightly greater mass in a nuclear reaction, the reactants or the products? Why?

It depends on the type of reaction. A fusion reaction will have greater mass in the products because it is combining two nuclei. A fission reaction will not change mass.

Nuclear Decay Key

AN Whitehead



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Table of Contents Nuclear Decay Key

1. Understanding the eBook Nuclear Decay Key
 - The Rise of Digital Reading Nuclear Decay Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Nuclear Decay Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Nuclear Decay Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Nuclear Decay Key
 - Personalized Recommendations

- Nuclear Decay Key User Reviews and Ratings
- Nuclear Decay Key and Bestseller Lists
- 5. Accessing Nuclear Decay Key Free and Paid eBooks
 - Nuclear Decay Key Public Domain eBooks
 - Nuclear Decay Key eBook Subscription Services
 - Nuclear Decay Key Budget-Friendly Options
- 6. Navigating Nuclear Decay Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Nuclear Decay Key Compatibility with Devices
 - Nuclear Decay Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Nuclear Decay Key
 - Highlighting and Note-Taking Nuclear Decay Key
 - Interactive Elements Nuclear Decay Key
- 8. Staying Engaged with Nuclear Decay Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Nuclear Decay Key
- 9. Balancing eBooks and Physical Books Nuclear Decay Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Nuclear Decay Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Nuclear Decay Key
 - Setting Reading Goals Nuclear Decay Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Nuclear Decay Key
 - Fact-Checking eBook Content of Nuclear Decay Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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