

Ex 1. For which of the following reactions would the ΔH° and the equilibrium constant K_{eq} be

1) $\Delta H^\circ_{rxn} < 0$ & $K_{eq} > 1$ \rightarrow exothermic

2) $\Delta H^\circ_{rxn} < 0$ & $K_{eq} < 1$ \rightarrow endothermic

3) $\Delta H^\circ_{rxn} > 0$ & $K_{eq} > 1$ \rightarrow exothermic

4) $\Delta H^\circ_{rxn} > 0$ & $K_{eq} < 1$ \rightarrow endothermic

Ex 2. The ΔH° for the reaction between acetylene and hydrogen is

the following: $\Delta H^\circ_{rxn} = -130 \text{ kJ/mol}$

$\text{C}_2\text{H}_2(\text{g}) + 2\text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$ $\Delta H^\circ_{rxn} = -130 \text{ kJ/mol}$

What is ΔH°_{rxn} for the reaction: $\text{C}_2\text{H}_2(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_4(\text{g})$ $\Delta H^\circ_{rxn} = ?$

Ex 3. Calculate the ΔH°_{rxn} for the reaction of H_2 and O_2 to form H_2O .

$\text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta H^\circ_{rxn} = -285.8 \text{ kJ/mol}$

$\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{l})$ $\Delta H^\circ_{rxn} = -571.6 \text{ kJ/mol}$

Ex 4. Calculate ΔH°_{rxn} for the reaction

Ex 5. Calculate ΔH°_{rxn} for the following reaction: $\text{C}_2\text{H}_2(\text{g}) + 2\text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$

$\text{C}_2\text{H}_2(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_4(\text{g})$ $\Delta H^\circ_{rxn} = -174.4 \text{ kJ/mol}$

$\text{C}_2\text{H}_4(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$ $\Delta H^\circ_{rxn} = -136.7 \text{ kJ/mol}$

Ex 6. Calculate ΔH°_{rxn} for the reaction: $\text{C}_2\text{H}_2(\text{g}) + 2\text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$

Ex 7. Calculate the ΔH°_{rxn} for the reaction of H_2 and O_2 to form H_2O and H_2O_2 .

$\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$

Ex 8. Calculate ΔH°_{rxn} for the reaction

$\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$

$\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}_2(\text{l})$

Ex 9. Calculate ΔH°_{rxn} for the reaction: $\text{C}_2\text{H}_2(\text{g}) + 2\text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$

Ex 10. Calculate the ΔH°_{rxn} for the reaction: $\text{C}_2\text{H}_2(\text{g}) + 2\text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$

$\text{C}_2\text{H}_2(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_4(\text{g})$ $\Delta H^\circ_{rxn} = -174.4 \text{ kJ/mol}$

$\text{C}_2\text{H}_4(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$ $\Delta H^\circ_{rxn} = -136.7 \text{ kJ/mol}$

$\text{C}_2\text{H}_2(\text{g}) + 2\text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$ $\Delta H^\circ_{rxn} = -211.1 \text{ kJ/mol}$

Ex 11. Calculate ΔH°_{rxn} for the reaction: $\text{C}_2\text{H}_2(\text{g}) + 2\text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$

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