Basic Chemistry Study Guide 2

Hydrogen bonds

- Form between two polar molecules covalent bonds
- A hydrogen's &+ region is attracted to a &- region in another molecule.
- Form between water molecules (Where are the &+ and &- regions in a water molecule?) hydrogen bonds, at the top of the symbol
- Each hydrogen bond is very weak (can be easily formed and broken).
- There are many hydrogen bonds in water. Together, these bonds give water its unique properties that are essential for life
- One of these properties is the ability to be a solvent for other polar molecules.
- Hydrogen bonds are depicted in figures as dotted or dashed lines between two atoms.

2. Solutions, solvents, and solutes

- Water is the most abundant molecule in living organisms, and 60% of human body weight is water.
- Water is located inside and outside of our cells.
- 3. Body fluids are solutions
- Define the terms solution, solvent, and solute. Solvent a substance capable of dissolving other polar molecules and ionic bonds, solutes, molecules dissolved in the solvent, solution homogeneous mixture.

3. How substances dissolve in water

- When a solute dissolves, it becomes surrounded by water molecules.
- 2. For a solute to dissolve, it must attract water molecules.
- What kinds of substances attract water molecules? Sugars, nucleic acids, and amino acids
- When a solute dissolves it forms hydrogen bonds with water molecules.
- Define the terms ion, cation, and anion. Ion atom with a charge, cation; positively charged ion, anions; negatively charged ion.
- Define the terms hydrophilic and hydrophobic Hydrophilic attracted to water, hydrophobic, a kaid of water
- What types of molecules are hydrophilic and what types are hydrophobic? Nonpolar hydrophobic, polar hydrophilic

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