

Part B. Balancing and Classifying Chemical Equations

#	Chemical Equation	Classification	
1	$\text{CH}_3\text{CH}_2\text{OH} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$		
2	$\text{C}_{12}\text{H}_{22}\text{O}_{11} \rightarrow \text{C} + \text{H}_2\text{O}$		
3	$\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$		
4	$\text{Cu} + \text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + \text{Ag}$		
5	$\text{H}_2\text{O}_2 \rightarrow \text{H}_2\text{O} + \text{O}_2$		
6	$\text{Zn} + \text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$		
7	$\text{HCl} + \text{Na}_2\text{CO}_3 \rightarrow \text{H}_2\text{CO}_3 + \text{NaCl}$		
8	$\text{CaCl}_2 + \text{Na}_3\text{PO}_4 \rightarrow \text{Ca}_3(\text{PO}_4)_2 + \text{NaCl}$		
9	$\text{FeCl}_3 + \text{KSCN} \rightarrow \text{Fe}(\text{SCN})_3 + \text{KCl}$		
10	$\text{H}_2\text{SO}_4 + \text{NaOH} \rightarrow \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$		

Study Guide Classifying Chemical Equations

JG Myers



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