AP Chemistry Final Study Guide

Unit 1: Atomic Structure and Properties (7%-9%)

- At STP, one mole of gas always occupies 22.4 L
- Mass Spectroscopy:
 - Atoms have different isotopes (carbon-12, carbon-14, etc) and this will affect the average atomic mass
 - The most abundant isotope will have the greatest effect
- Coulomb's Law:
 - The greater the charge of the nucleus, the more energy is required to ionize an electron
 - The closer an electron is to the nucleus, the more energy is required to ionize an electron
- · Electromagnetic Radiation:
 - When atoms absorb energy, electrons jump to higher energy levels
 - When electrons drop from higher to lower energy levels, atoms give off electromagnetic radiation energy
- Kinetic Energy of Ionized Electrons:
 - The faster an ejected electron is going, the more kinetic energy it has
 - Electrons originally further away from the nucleus require less energy to eject, so
 it will have more kinetic energy
- · Electron Configuration:
 - When entering the d-block (transition metals), the energy level drops by 1 unless you are using shorthand notation
 - Nickel: 1s²2s²2p⁶3s²3p⁶3d⁸4s²(notice the d-block)
 - Nickel (with shorthand): [Ar]4s²3d⁸
- Law of Definite Proportions:
 - A pure chemical compound broken down into elements always contains elements of a fixed ratio, independent of where and how it was created
- Law of Multiple Proportions:
 - When elements form compounds, the proportions of the elements in those chemical compounds can be expressed in small whole number ratios
- Aufbau Principle:
 - When building up the electron configuration of an atom, electrons are placed in orbitals, subshells, and shells in order of increasing energy

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