

Pre Public Examination
GCSE Mathematics (Edexcel style)
May 2017
Higher Tier
Paper 2H

Name

Class

TIME ALLOWED

1 hour 30 minutes

INSTRUCTIONS TO CANDIDATES

- Answer **all** the questions.
- Read each question carefully. Make sure you know what you have to do before starting your answer.
- **You are permitted to use a calculator in this paper.**
- Do all rough work in this book.

INFORMATION FOR CANDIDATES

- The number of marks is given in brackets () at the end of each question or part question on the Question Paper.
- **You are reminded of the need for clear presentation in your answers.**
- The total number of marks for this paper is **80**.

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Question	Mark	Out of
1		2
2		4
3		5
4		2
5		2
6		4
7		4
8		4
9		5
10		4
11		5
12		4
13		3
14		4
15		4
16		4
17		4
18		4
19		6
20		6
Total		80

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Ying-Ying Zheng



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Table of Contents Ppe Pixl May 2015 Paper

1. Understanding the eBook Ppe Pixl May 2015 Paper
 - The Rise of Digital Reading Ppe Pixl May 2015 Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Ppe Pixl May 2015 Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ppe Pixl May 2015 Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ppe Pixl May 2015 Paper
 - Personalized Recommendations
 - Ppe Pixl May 2015 Paper User Reviews and Ratings
 - Ppe Pixl May 2015 Paper and Bestseller Lists
5. Accessing Ppe Pixl May 2015 Paper Free and Paid eBooks
 - Ppe Pixl May 2015 Paper Public Domain eBooks
 - Ppe Pixl May 2015 Paper eBook Subscription Services

- Ppe Pixl May 2015 Paper Budget-Friendly Options
- 6. Navigating Ppe Pixl May 2015 Paper eBook Formats
 - ePub, PDF, MOBI, and More
 - Ppe Pixl May 2015 Paper Compatibility with Devices
 - Ppe Pixl May 2015 Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ppe Pixl May 2015 Paper
 - Highlighting and Note-Taking Ppe Pixl May 2015 Paper
 - Interactive Elements Ppe Pixl May 2015 Paper
- 8. Staying Engaged with Ppe Pixl May 2015 Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ppe Pixl May 2015 Paper
- 9. Balancing eBooks and Physical Books Ppe Pixl May 2015 Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ppe Pixl May 2015 Paper
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ppe Pixl May 2015 Paper
 - Setting Reading Goals Ppe Pixl May 2015 Paper
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ppe Pixl May 2015 Paper
 - Fact-Checking eBook Content of Ppe Pixl May 2015 Paper
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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web $m \times c \times \Delta t$ equation 1 where m is the total mass of the solution solute plus solvent c is the specific heat of the solution and Δt is the observed temperature change the specific heat of the solution is generally assumed to be the same as that of
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web the heat of reaction ΔH_{soln} is written after the products in units of kJ mol rxn for example sodium hydroxide dissolves

exothermically releasing 44.2 kilojoules per mole dissolved this process is represented as $\text{NaOH (s)} \rightarrow \text{NaOH (aq)}$ $\Delta H_{\text{soln}} = -44.2 \text{ kJ/mol}$ rxn write equations to similarly represent the dissolving process for

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web hand warmers are usually placed inside the gloves to keep your hands warm the heat keeping your hands warm is formed as a result of an exothermic reaction which occurs in a small packet hand warmers have 5 main ingredients in them salt carbon polypropylene vermiculite water and iron

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web safety 1 all chemicals in the laboratory are to be considered dangerous do not touch taste or smell any chemicals unless specifically instructed to do so 2 never remove chemicals or other materials from the laboratory area 3 do not eat food drink beverages or chew gum in the laboratory

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web background the hand warmers are filled with sodium acetate click the metal disc and you will notice the liquid changes state and becomes a solid in the process heat is released the sodium acetate is a supersaturated solution

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web the goal of my experiment and research is to evaluate the various number of compounds to be used as hand warmers and locate the ones that can raise the temperature to 20 degrees celsius in 100 g water and costs as little as possible

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web in this lab you are challenged to use chemistry to determine which ionic compound makes the best hand warmer in this experiment you will collect data that will allow you to calculate the heat of solution ΔH_{soln} or the amount of energy released

or absorbed in kJ mol of the substance being dissolved

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web mar 4 2014 contains water the hand warmer is struck in a manner that ruptures the inner pouch releasing the ionic salt into the water of the outer pouch the salt dissolves and the water warms this type of hand warmer tends to produce a more vigorous heat than the dry powder type of hand warmer but does not produce heat for quite as long

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