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Ions in Aqueous Solutions and Colligative Properties

Section Quiz: Colligative Properties of Solutions

Write each answer provided within the circle of the quiz or phrase that best completes each statement or best answers each question.

- _____ 1. Which is/are not colligative properties of solutions?
- (a) It is directly proportional to the number of particles in solution.
 - (b) It is directly proportional to the number of particles in solution.
 - (c) It depends on the properties of an individual solute molecule.
 - (d) None of the above.
- _____ 2. Compounds with a total of 0.100 M CaCl_2 , $\text{Al}_2(\text{SO}_4)_3$, and NaCl solutions give _____.
- (a) the same freezing-point depression.
 - (b) none of the following gives depression.
 - (c) the same freezing-point elevation.
 - (d) none are none the freezing-point depression.
- _____ 3. Compounds with a total of 0.100 M NaCl , CaCl_2 , $\text{Al}_2(\text{SO}_4)_3$, and Na_2SO_4 solutions _____.
- (a) the same freezing-point depression.
 - (b) all of the above the freezing-point depression.
 - (c) a freezing-point depression about two-thirds of (b).
 - (d) a freezing-point depression about 1/3 of (b) or (c).
- _____ 4. When a nonelectrolyte solute dissolved in water, the liquid portion of the solvent _____.
- (a) increases.
 - (b) decreases.
 - (c) stays the same.
 - (d) changes depending on the solute used.
- _____ 5. Which of the following compounds would be most effective in lowering the boiling point of 1.0 M water?
- (a) CaCl_2
 - (b) NaCl
 - (c) $\text{Al}_2(\text{SO}_4)_3$
 - (d) Na_2SO_4

Rosengarten Colligative Properties Homework Answers

L Cohen



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