

Find the temperature  $T$  at which  $x_A + x_B = 1.0$

By trial,  $T = 91^\circ\text{C}$

$$x_A = 152 / 198 = 0.768$$

$$x_B = 95 / 409.5 = 0.232$$

At  $91^\circ\text{C}$ ,  $P_w^*$  is about 530 mm Hg, well above  $p_w$ . Hence no water will condense at  $87^\circ\text{C}$ .

- (c) Basis: 1 mole of feed (0.35 moles water, 0.65 mol organic). Must assume that some liquid water will be present, exerting its own vapor pressure at whatever the temperature is. Find this temperature by trial.

By trial  $71.8^\circ\text{C}$ .  $p_w = 250$  mm Hg

For the organic,  $P^* = 380 - 250 = 130$  mm Hg

Moles of water in vapor:  $(250 / 380) \times 0.5 = 0.329$

Moles of organic in vapor:  $0.5 - 0.329 = 0.171$

Fraction of the organic vaporized:

$$f = 0.171 / 0.65 = 0.271$$

Composition of organic liquid fed:

$$x'_{FA} = 0.40 / 0.65 = 0.6154$$

$$x'_{FB} = 0.25 / 0.65 = 0.3846$$

From Eq. (18.2),

$$y_B = -\frac{1-f}{f} x'_B + \frac{x'_{FB}}{f} = 1.463 - 2.802 x'_B$$

From the equations before Eq. (21.39),

$$y_B = P'_B x'_B / 130$$

$$1.463 - 2.802 x'_B = P'_B x'_B / 130$$

$$x'_B = \frac{1.463}{2.802 + (P'_B / 130)}$$

# Solution Manual Kinetics Harriott

**Warren Lee McCabe, Julian Cleveland  
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