

Experiment No. 06

Date:

Name of the experiment: Determination of pKa value of weak acid.

Principle

A buffer solution is one which resists (or buffers) a change in its pH. If acid is added then, within reason, the pH does not fall; if base is added, the pH does not rise. Buffers are usually composed of a mixture of weak acids or weak bases and their salts and function best at a pH equal to the pKa of the acid or base involved in the buffer. The equation that predicts the behaviour of buffers is known as the Henderson-Hasselbalch equation (named after chemists Lawrence Joseph Henderson and Karl Albert Hasselbalch). It is derived as follows, by considering a weak acid that ionises in solution.

$$\text{pH} = \text{pK}_a + \log \frac{[\text{Salt}]}{[\text{Acid}]}$$

$$\text{or } \text{pK}_a = \text{pH} - \log \frac{[\text{Salt}]}{[\text{Acid}]}$$

Where, $\text{pK}_a = -\log K_a$ and K_a = Dissociation constant of acid.

An example of a buffer is a mixture of acetic acid and sodium acetate, which ionises and acts as follows:



Apparatus:

1. Electrical balance
2. Spatula
3. Volumetric flask
4. Measuring cylinder
5. Burette
6. Pipette and pipette filler
7. Conical flask
8. Beaker
9. Funnel

Chemicals / Reagents:

1. 0.1 N Sodium hydroxide (NaOH) solution
2. 1% Acetic acid (CH_3COOH) solution
3. 0.1 N Oxalic acid ($\text{H}_2\text{C}_2\text{O}_4$) solution
4. 0.5% Phenolphthalein indicator
5. Distilled Water

Calculations: Full calculations required to make the reagents, for the experiment must be written down.

Preparation of Reagents:

A. 1N 100 mL Acetic Acid Solution: Take 1 mL glacial acetic acid (100%) in a 100 mL volumetric flask, and then adjust the volume up to 100 mL mark by adding distilled water.

B. 0.1 N 1000 mL NaOH Solution: Completely dissolve 4.12 gm NaOH (97%) pellets in a 1000 mL volumetric flask at first with a small portion of DW, and finally adjust the volume up to the mark by adding DW Q.S.

C. 0.1 N 100 mL Oxalic Acid Solution: Completely dissolve 0.636 gm oxalic acid (99%) in a 100 mL volumetric flask at first with a small portion of DW, and finally make the volume up to mark by adding distilled water Q.S.

D. 0.5% Phenolphthalein: 50 mL Ethanol + 0.5 gm phenolphthalein + 50 mL distilled water.

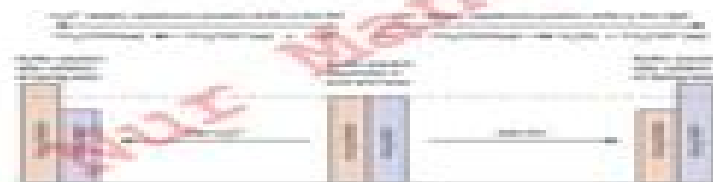


Fig. Mechanism of Buffer action of an acid buffer.



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Robert A. Alberty, Robert J. Silbey



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Scientific and Technical Books and Serials in Print, 1984 *Synthesis and Technique in Inorganic Chemistry* Gregory S. Girolami, Thomas B. Rauchfuss, Robert J. Angelici, 1999 Previously by Angelici this laboratory manual for an upper level undergraduate or graduate course in inorganic synthesis has for many years been the standard in the field In this newly revised third edition the manual has been extensively updated to reflect new developments in inorganic chemistry Twenty three experiments are divided into five sections solid state chemistry main group chemistry coordination chemistry organometallic chemistry and bioinorganic chemistry The included experiments are safe have been thoroughly tested to ensure reproducibility are illustrative of modern issues in inorganic chemistry and are capable of being performed in one or two laboratory periods of three or four hours Because facilities vary from school to school the authors have included a broad range of experiments to help provide a meaningful course in almost any academic setting Each clearly written illustrated experiment begins with an introduction that highlights the theme of the experiment often including a discussion of a particular characterization method that will be used followed by the experimental procedure a set of problems a listing of suggested Independent Studies and literature references

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