

# solutions, acids, and bases

## 1. VOCABULARY REVIEW OF SOLUTIONS

**Solute** - what is **dissolved** in the solvent to form a solution.

**Solvent** - the substance that **dissolves** the solute, ordinarily a liquid when the solute and the solvent come together to form a **homogeneous mixture**.

**miscible** - two mix in all proportions (none to be mixed) - liquids will pour each over

**immiscible** - don't mix in any proportions (none to be mixed)

examples of miscible/immiscible - ethanol/water/methanol, etc. and oil/water

**solvent** - capable of being dissolved in liquid

**insoluble** - a substance (tends to be a solid) that **will not dissolve** in a solvent - an example is sand and water

**suspension** - looks like a solution when **stirred**, the **solutes** a bit of water (suspensions) also it is a mixture - gaseous mixture where the small solid particles of a substance are spread throughout a liquid without settling.

**undersaturated** - solutes with solubility chart when the temperature would be enough to make a substance dissolve but at a greater amount of solute solubility **less solute dissolves than can** (ex. 100g @ 20°C + 100g)

**saturated** - solutes with solubility chart when the temperature - solubility relationship is at the line in other words a solution with solute has **reached** at the maximum possible amount.

**supersaturated** - contains **more than the max.** amount of solute that can be dissolved at the temperature on the solubility chart this would mean the grams of solute would go above the line.

## 2. SOLUBILITY VS. TEMPERATURE CHART



above the line - **supersaturated**

at the line - **saturated**

under the line - **undersaturated**

## 3. METHODS OF SEPARATION

**filtration** - separate particles by **size**

**distillation** - separate by **boiling point** - where two liquids to be separated.

**evaporation** - where one to remove a **liquid from a solid** by **transforming** the liquid to a **gaseous state**.

**centrifugation** - separating molecules by their densities by spinning them apart.

**decanting** - process of separating insoluble or suspension solutions by **pouring** away the liquid from the solid.

**chromatography** - separation by the **rates** at which components of different solutions move through a medium (ex. gel electrophoresis)

# Solutions Acids And Bases Study Guide Key

**David R. Klein, Laurie S. Starkey**



## **Solutions Acids And Bases Study Guide Key:**

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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