

# Test Solutions and Solubility SCH3U

Name: \_\_\_\_\_

15 K   14 T   9 C   10 A   Assessment of Learning (Product)

Multiple Choice. Circle the most correct answer. (15K, 5T)

1. How many of the following salts can be expected to be insoluble in water?

sodium sulfate   potassium chloride   ammonium nitrate   potassium chromate

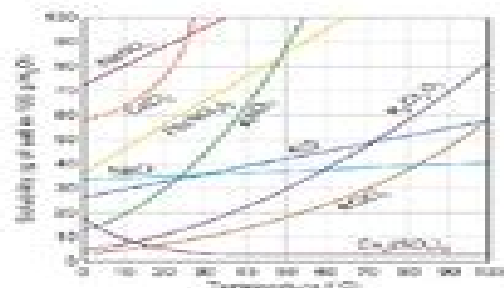
A) two   B) three   C) one   D) four   E) none

2. When solutions of silver nitrate and potassium chloride are mixed, which of the following will be present in the net ionic equation?

A)  $\text{KNO}_3(\text{aq})$    B)  $\text{K}^+$    C)  $\text{KCl}(\text{aq})$    D)  $\text{Cl}^-$    E) None of these

3. The net ionic equation for formation of an aqueous solution of  $\text{NiI}_2$  accompanied by evolution of  $\text{CO}_2$  gas via mixing solid  $\text{NiCO}_3$  and aqueous hydroiodic acid is \_\_\_\_\_.

- A)  $2\text{NiCO}_3(\text{s}) + \text{HI}(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g}) + 2\text{Ni}^{2+}(\text{aq})$   
 B)  $\text{NiCO}_3(\text{s}) + \text{I}^-(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g}) + \text{Ni}^{2+}(\text{aq}) + \text{HI}(\text{aq})$   
 C)  $\text{NiCO}_3(\text{s}) + 2\text{H}^+(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g}) + \text{Ni}^{2+}(\text{aq})$   
 D)  $\text{NiCO}_3(\text{s}) + 2\text{HI}(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g}) + \text{NiI}_2(\text{aq})$   
 E)  $\text{NiCO}_3(\text{s}) + 2\text{HI}(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g}) + \text{Ni}^{2+}(\text{aq}) + 2\text{I}^-(\text{aq})$



4. A beaker containing 80 grams of lead(II) nitrate,  $\text{Pb}(\text{NO}_3)_2$ , in 100 grams of water has a temperature of 30 °C. Approximately how many grams of the salt are undissolved, on the bottom of the beaker?

A) 14 grams   B) 66 grams   C) 80 grams   D) 20 grams

5. A solution in which no more solid can dissolve is called a \_\_\_\_\_ solution:

- A) solute   B) solution   C) electrolyte   D) immiscible   E) concentrated  
 F) soluble   G) solvent   H) unsaturated   I) dilute   J) non-electrolyte  
 K) saturated   L) miscible

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**Carl Remigius Fresenius**



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