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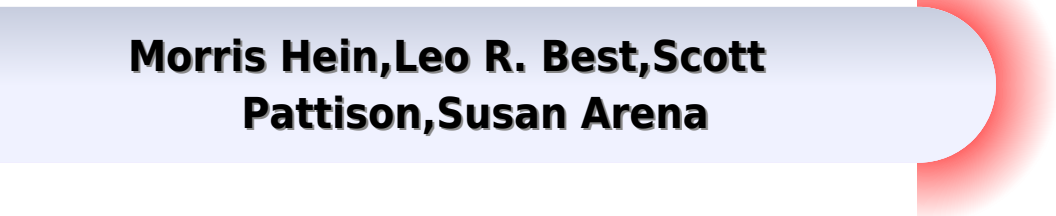
End of chapter answers

Chapter 2

1. a) The properties are similar because both gas and liquid are fluids.
 b) All phase transitions of water are unusual. The liquid to gas because it occurs at a high temperature, due to hydrogen bonding. The liquid-solid transition is unusual because the maximal density of water occurs at 4°C, due to closer approach of water molecules in the liquid.
 c) Most of the unusual properties are due to the extensive hydrogen bonding possible because of the OH bond and the geometry of the molecule, forming extensive three dimensional lattices.
2. It is possible. Using the density (1 g/ml) and molecular weight (18 g/mol) of water, the concentration is 1/18 mol/l or 55.5 M.
3. Electronegativities are calculated from molecules in which the bonds occur. They are different in different molecules due to other forces (through bond and through space electrical effects), so the values are averaged. The scale for each atom is thus a relative one; the difference between atoms can be easily estimated by subtracting the electronegativity values of each. The 1-4 scale was arbitrarily chosen.
4. Sulfur is larger, the electrons more diffuse, and as a result is unable to form a hydrogen bond.
5. They are in adjacent columns in the periodic table, so we can expect some properties to be shared, but have distinct valences. The nitrogen atom has five electrons, $1s^2 2p^3$. The p orbital has 6 potential electrons; the 3 in nitrogen atom are thus half-filled. Once bound to other atoms, the electrons are hybrids, such as the common hydride, NH_3 . The extra electron pair in this molecule makes it a Lewis base (electron pair donor). However, N can occur in acids too, such as in nitric acid. As amines, the reason for the tendency to positive charge is that when amines accept a proton (Bronstead base) they become positively charged.
 The oxygen atom has six electrons, $1s^2 2p^4$. It also hybridizes in forming compounds; in this case many of its compounds become acids such as $R-COOH$. When this loses a proton, it becomes negatively charged. However, the oxygen hydride analogous to $NH_3 - H_2O$ - can be considered both basic and acidic.
6. CO_2 has polar bonds, but is not a polar molecule because the polarity vectors cancel out. The geometry of the molecule is linear.
7. NH_4^+ dissolves in water because it has a full charge, which attracts the partial negative charges of water even more powerfully than another polar molecule does.

Solutions Manual Biochemistry

**Morris Hein, Leo R. Best, Scott
Pattison, Susan Arena**



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2 2 0 20 0 2 10 1 9 msb di atas diimplementasikan dalam r sebagai berikut ini persamaan

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