

not significant and zeros at the end of a number and after the decimal point are always significant.

(c) 18 students has an infinite number of significant figures since this is an exact number.

- 1.10 (a) Since the digit to be dropped (the second 4) is less than 5, round down. The result is 3.774 L.

(b) Since the digit to be dropped (0) is less than 5, round down. The result is 255 K.

(c) Since the digit to be dropped is equal to 5 with nothing following and the preceding digit is even, round down. The result is 55.26 kg.

- 1.11 (a)
$$\begin{array}{r} 24.567 \text{ g} \\ + 0.04478 \text{ g} \\ \hline 24.61178 \text{ g} \end{array}$$
 This result should be expressed with 3 decimal places. Since the digit to be dropped (7) is greater than 5, round up. The result is 24.612 g (5 significant figures).

(b) $4.6742 \text{ g} / 0.00371 \text{ L} = 1259.89 \text{ g/L}$

0.00371 has only 3 significant figures so the result of the division should have only 3 significant figures. Since the digit to be dropped (first 9) is greater than 5, round up. The result is 1260 g/L (3 significant figures).

(c)
$$\begin{array}{r} 0.378 \text{ mL} \\ + 42.3 \text{ mL} \\ - 1.5833 \text{ mL} \\ \hline 41.0947 \text{ mL} \end{array}$$
 This result should be expressed with 1 decimal place. Since the digit to be dropped (9) is greater than 5, round up. The result is 41.1 mL (3 significant figures).

- 1.12 (a) Estimate: $^{\circ}\text{F} = 2 \times ^{\circ}\text{C}$ if $^{\circ}\text{C}$ is large. The melting point of gold = 2000 $^{\circ}\text{F}$.

$$\text{Calculation: } ^{\circ}\text{F} = \left(\frac{9}{5} \times ^{\circ}\text{C}\right) + 32 = \left(\frac{9}{5} \times 1064\right) + 32 = 1947^{\circ}\text{F}$$

(b) $r = d/2 = 3 \times 10^{-6} \text{ m} = 3 \times 10^{-4} \text{ cm}$; $h = 2 \times 10^{-6} \text{ m} = 2 \times 10^{-4} \text{ cm}$

$$\text{Estimate: volume} = \pi r^2 h = 3r^2 h = 3(3 \times 10^{-4} \text{ cm})^2 (2 \times 10^{-4} \text{ cm}) = 5 \times 10^{-11} \text{ cm}^3$$

$$\text{Calculation: volume} = \pi r^2 h = (3.1416)(3 \times 10^{-4} \text{ cm})^2 (2 \times 10^{-4} \text{ cm}) = 6 \times 10^{-11} \text{ cm}^3$$

- 1.13 1 mi = 1760 yd and 1 m = 1.0936 yd

$$26 \text{ mi} \times \frac{1760 \text{ yd}}{1 \text{ mi}} = 45,760 \text{ yd}$$

$$\text{marathon in yards} = 45,760 \text{ yd} + 385 \text{ yd} = 46,145 \text{ yd}$$

$$\text{marathon in meters} = 46,145 \text{ yd} \times \frac{1 \text{ m}}{1.0936 \text{ yd}} = 42,196 \text{ m}$$

- 1.14 1 carat = 200 mg = $200 \times 10^{-3} \text{ g} = 0.200 \text{ g}$

$$\text{Mass of Hope Diamond in grams} = 44.4 \text{ carats} \times \frac{0.200 \text{ g}}{1 \text{ carat}} = 8.88 \text{ g}$$

$$1 \text{ ounce} = 28.35 \text{ g}$$

$$\text{Mass of Hope Diamond in ounces} = 8.88 \text{ g} \times \frac{1 \text{ ounce}}{28.35 \text{ g}} = 0.313 \text{ ounces}$$

Solutions Manual Chemistry Assessment

**Andrew Burrows, John Holman, Simon
Lancaster, Andrew Parsons, Tina
Overton, Gwen Pilling, Gareth Price**



Solutions Manual Chemistry Assessment:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter. General Organic and Biological Chemistry Kenneth W.

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