

# AP Chemistry - Chemical Kinetics

Name: \_\_\_\_\_ Date: \_\_\_\_\_ Period: \_\_\_\_\_

- After five half-life periods for a first order reaction, what fraction of reactant remains?
- If you plot  $\ln[\text{reactant}]$  versus time and observe a straight line, what is the order of the reaction? If  $1/[\text{reactant}]$  is plotted versus time, and a straight line of negative slope is observed, what is the order of the reaction?
- Indicate whether each of the following statements is true or false. Change the wording of each false statement to make it true.
  - In stepwise reactions the rate-determining step is the slow one.
  - It is possible to change the rate constant for a reaction by changing the temperature.
  - As a first order reaction proceeds at a constant temperature, the rate remains constant.
  - The rate constant for a reaction is independent of reactant concentrations.
- For each reaction below give the relative rates of formation of products and disappearance of reactants using coefficients and rate change symbols.
  - $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
  - $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$
  - $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$
- The reaction  $\text{CO}_2 + \text{NO}_2 \rightarrow \text{CO}_2 + \text{NO}_2$  is second order in  $\text{NO}_2$  and zero order in  $\text{CO}$  at temperatures less than 500K.
  - Write the rate law for the reaction.
  - How will the reaction rate change if the  $\text{NO}_2$  concentration is halved?
- For a reaction of the type  $\text{A} + \text{B} \rightarrow \text{C}$ , it is experimentally observed that doubling the concentration of B causes the reaction rate to be increased fourfold, but doubling the concentration of A has no effect on the rate. Determine the initial rate equation.
- A study of the reaction  $2\text{A} + \text{B} \rightarrow \text{C} + \text{D}$  gave the following results:

Experiment	Init. Con. (M)		Init. Rate (M/s)
	$[\text{A}]_0$	$[\text{B}]_0$	
1	0.10	0.10	$6.0 \times 10^{-3}$
2	0.20	0.10	$1.2 \times 10^{-2}$
3	0.10	0.20	$1.2 \times 10^{-2}$
4	0.20	0.20	$1.2 \times 10^{-2}$

- What is the rate law for this reaction?
  - What is the overall order of the reaction?
  - Could this reaction occur in a single step?
- The decomposition of  $\text{N}_2\text{O}_5$  (to give  $\text{NO}_2$  and  $\text{O}_2$ ) follows the rate law "rate =  $k[\text{N}_2\text{O}_5]$ " where  $k$  is  $5.0 \times 10^{-3} \text{ s}^{-1}$  per second at a particular temperature.
    - What is the half-life of this reaction?
    - How long does it take for the  $\text{N}_2\text{O}_5$  concentration to drop to one tenth of its original value?

# Solutions Manual For Ap Chem Chemical Kinetics

**Rachel S Tattersall**



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