

MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question.

- When $\text{Mg}(\text{OH})_2$ reacts with H_2SO_4 what are the products?
 A) MgSO_4 and H_2O B) MgO and H_2SO_4 C) $\text{Mg}(\text{OH})_2$ and H_2O
- Which of the following is a strong acid?
 A) CaSO_4 B) HNO_3 C) HF
- Which of the following is a precipitation reaction?
 A) $\text{NaI}(\text{aq}) + \text{AgNO}_3(\text{aq}) \rightarrow \text{AgI}(\text{s}) + \text{NaNO}_3(\text{aq})$
 B) $\text{HBr}(\text{aq}) + \text{LiOH}(\text{aq}) \rightarrow \text{LiBr}(\text{aq}) + \text{H}_2\text{O}(\text{l})$
 C) $\text{Cu}(\text{s}) + 2 \text{AgNO}_3(\text{aq}) \rightarrow 2 \text{Ag}(\text{s}) + \text{Cu}(\text{NO}_3)_2(\text{aq})$
 D) $2 \text{HCl}(\text{aq}) + \text{Na}_2\text{SO}_4(\text{s}) \rightarrow 2 \text{NaCl}(\text{aq}) + \text{H}_2\text{O}(\text{l}) + \text{SO}_4(\text{g})$
- Ozone, O_3 , is a product in automobile exhaust by the equation $\text{NO}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{NO}(\text{g}) + \text{O}_3(\text{g})$. What mass of ozone is predicted to form from the reaction of 2.0 g NO_2 in a car's exhaust and excess oxygen?
 A) 1.8 g O_3 B) 1.1 g O_3 C) 2.1 g O_3
- What precipitate forms when mixing silver(I) nitrate AgNO_3 and sodium chloride NaCl ?
 A) silver(I) chloride B) sodium bromide C) sodium nitrate
- Which of the following reactions are written correctly?
 A) $\text{H}_2\text{NO}_3(\text{aq}) + 2 \text{NaOH}(\text{aq}) \rightarrow \text{Na}_2\text{NO}_3(\text{aq}) + 2 \text{H}_2\text{O}(\text{l})$
 B) $\text{HNO}_3(\text{aq}) + \text{NaOH}(\text{aq}) \rightarrow \text{NaNO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l})$
 C) $\text{HNO}_3(\text{aq}) + \text{K}_2\text{OH}(\text{aq}) \rightarrow \text{K}_2\text{NO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l})$
- The reaction $\text{HNO}_3(\text{aq}) + \text{NaOH}(\text{aq}) \rightarrow \text{NaNO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l})$ is an example of _____.
 A) an acid-base reaction B) a precipitation reaction C) a combustion reaction
- Which is the correct balanced reaction for the combustion of octane?
 A) $\text{C}_8\text{H}_{18}(\text{g}) + 12.5 \text{O}_2(\text{g}) \rightarrow 8 \text{CO}_2(\text{g}) + 9 \text{H}_2\text{O}(\text{g})$
 B) $\text{C}_8\text{H}_{18}(\text{g}) + 25 \text{O}(\text{g}) \rightarrow 8 \text{CO}_2(\text{g}) + 9 \text{H}_2\text{O}(\text{g})$
 C) $2 \text{C}_8\text{H}_{18}(\text{g}) + 25 \text{O}_2(\text{g}) \rightarrow 16 \text{CO}_2(\text{g}) + 18 \text{H}_2\text{O}(\text{g})$
- How many moles of HCl are needed to completely react with 50.0 mL of 0.250 M NaOH solution?
 A) 1.25×10^1 moles B) 1.25×10^{-2} moles C) 2.50×10^{-1} moles
- What is the oxidation number of Cr in a dichromate anion $\text{Cr}_2\text{O}_7^{2-}$?
 A) +6 B) +2 C) +3
- Write the net ionic equation for a reaction between a strong acid and strong base.
 A) $\text{H}_2\text{O}(\text{l}) \rightarrow \text{H}^+(\text{aq}) + \text{OH}^-(\text{aq})$
 B) $\text{H}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$
 C) $\text{HCl}(\text{aq}) + \text{NaOH}(\text{aq}) \rightarrow \text{NaCl}(\text{aq}) + \text{H}_2\text{O}(\text{l})$

Solution Stoichiometry Multiple Choice Answer

Paul Foglino, Princeton Review (Firm)



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