









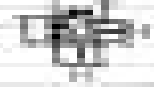



# Molecular Geometry

How can molecular shapes be predicted using the VSEPR theory?

## Why?

When you draw a Lewis structure for a molecule on paper, you are making a two-dimensional representation of the atoms. In reality, however, molecules are not flat—they are three-dimensional. The 3D shape of a molecule is important because it determines many physical and chemical properties for the substance. In this activity you will learn how to predict molecular shapes.

## Model 1 – Lewis Structures

| Lewis Structures |   | 3-D Molecular Shape  |   |
|------------------|---|--|---|
| 1                | $\text{H}_2\text{O}$<br>     | $\text{H}_2\text{O}$<br>2 electron domains<br>(2 bonding, 2 nonbonding)    |    |
| 2                | $\text{CO}_2$<br>            | $\text{CO}_2$<br>2 electron domains<br>(2 bonding, 0 nonbonding)           |    |
| 3                | $\text{CH}_4$<br>            | $\text{CH}_4$<br>4 electron domains<br>(4 bonding, 0 nonbonding)           |    |
| 4                | $\text{NH}_3$<br>            | $\text{NH}_3$<br>4 electron domains<br>(3 bonding, 1 nonbonding)           |    |
| 5                | $\text{H}_2\text{SO}_4$<br> | $\text{H}_2\text{SO}_4$<br>4 electron domains<br>(4 bonding, 0 nonbonding) |   |
| 6                | $\text{SO}_2$<br>          | $\text{SO}_2$<br>3 electron domains<br>(2 bonding, 1 nonbonding)           |  |

Source: <http://www.khanacademy.org>

# Pogil Activities For High School Molecular Geometries Answers

**Jenn-Tai Hwang**



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