

Temperature, Thermal Energy, and Heat

Understanding Main Ideas

Answer the following questions in the space provided.

1. Explain how to distinguish the kinetic thermal energy. Relating these to a microscopic picture.

Ans. Kinetic energy is the energy of motion. It is the energy of the particles in the system. It is the energy of the particles in the system. It is the energy of the particles in the system.

2. Describe what is meant by absolute zero.

Ans. Absolute zero is the temperature at which the particles in the system have no kinetic energy. It is the temperature at which the particles in the system have no kinetic energy. It is the temperature at which the particles in the system have no kinetic energy.

3. Distinguish among the concepts of temperature, thermal energy, and heat.

Ans. Temperature is a measure of the average kinetic energy of the particles in the system. Thermal energy is the energy of the particles in the system. Heat is the energy that is transferred from one system to another.

4. Compare the following units to make up the Celsius and the Fahrenheit temperature scales.

Ans. The Celsius scale is based on the freezing and boiling points of water. The Fahrenheit scale is based on the freezing and boiling points of water.

5. Compare the changes in temperature of 1°C and 1°F. Is 1°C a larger change than 1°F?

Ans. 1°C is a larger change than 1°F. 1°C is equal to 1.8°F.

Building Vocabulary

Match each term with its definition, using the letter indicated in parentheses. Write your answers in the boxes that follow in the left column.

16. a heating

17. b temperature

18. c heat

19. d absolute zero

- a. the energy that is transferred from one system to another

- b. a measure of the average kinetic energy of the particles in the system

- c. the transfer of thermal energy from one system to another

- d. the temperature at which the particles in the system have no kinetic energy

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

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