



## End Semester Exam – Spring 2023 Programme: BE (Mechanical Engineering)

Subject: *Thermodynamics-II*

Course Code: *ME-230*

Total Marks: *100*

Time Allowed: *3 Hours*

Student Name: \_\_\_\_\_

Registration # \_\_\_\_\_

Section: \_\_\_\_\_

Signature: \_\_\_\_\_

### Instructions:

1. This is a closed book examination.
2. Only basic calculators are allowed. Digital diaries / programmable calculators are not allowed.
3. Question paper must be submitted with the answer book at the end of the exam.

### Q-1 (CLO 3):

A vertical piston–cylinder device initially contains 0.1 m<sup>3</sup> of helium at 20°C. The mass of the piston is such that it maintains a constant pressure of 300 kPa inside. A valve is now opened, and helium is allowed to escape until the volume inside the cylinder is decreased by one-half. Heat transfer takes place between the helium and its surroundings at 20°C and 95 kPa so that the temperature of helium in the cylinder remains constant.

(The gas constant of helium is  $R = 2.0769 \text{ kPa}\cdot\text{m}^3/\text{kg}\cdot\text{K}$ . The specific heats of helium are  $C_p = 5.1926 \text{ kJ/kg}\cdot\text{K}$  and  $C_v = 3.1156 \text{ kJ/kg}\cdot\text{K}$ )

**Interpret and Analyze the exergy principles involved in this problem by finding the following**

- The maximum work potential of the helium at the initial state (10)
- The exergy destroyed during this process. (10)

(Marks 20)

### Solution

From the ideal gas relation, the initial and the final masses in the cylinder are determined to be

$$m_1 = \frac{P_1 V}{RT_1} = \frac{(300 \text{ kPa})(0.1 \text{ m}^3)}{(2.0769 \text{ kPa} \frac{\text{m}^3}{\text{kg}\cdot\text{K}})(293 \text{ K})} = 0.0493 \text{ kg} \quad \text{.....(1)}$$

$$m_e = m_2 = \frac{m_1}{2} = 0.0247 \text{ kg} \quad \text{.....(1)}$$

The work potential of helium at the initial state is simply the initial exergy of helium, and is determined from the closed-system exergy relation

$$\Phi_1 = m_1 \phi = m_1 [(u_1 - u_o) - T_o(s_1 - s_o) + P_o(v_1 - v_o)] \quad \text{.....(1)}$$

$$v_1 = \frac{RT_1}{P_1} = \frac{(2.0769)(293)}{300} = 2.0284 \text{ m}^3/\text{kg} \quad \text{.....(1)}$$

$$v_o = \frac{RT_o}{P_o} = \frac{(2.0769)(293)}{95} = 6.405 \text{ m}^3/\text{kg} \quad \text{.....(1)}$$

And

$$s_1 - s_o = C_p \ln \frac{T_1}{T_o} - R \ln \frac{P_1}{P_o} = -2.388 \frac{\text{kJ}}{\text{kg}\cdot\text{K}} \quad \text{.....(1)}$$

Putting the values,

$$\Phi_1 = m_1 \phi = (0.0493)[\{3.1156(20 - 20)\} - (293)(-2.388) + (95)(2.0284 - 6.405)] = 14.0 \text{ kJ} \quad \text{.....(4)}$$

# Thermodynamics Final Exam

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