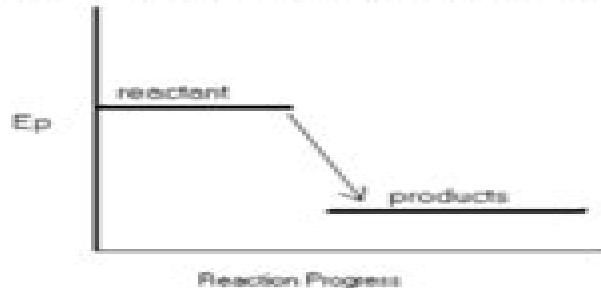


Thermodynamics and Rates of Reactions

Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

1. An endothermic reaction is one where
 - a. heat is transferred from the surroundings into a system
 - b. heat is transferred from a system into the surroundings
 - c. kinetic energy is transformed into potential energy
 - d. there is no transfer of heat
 - e. none of the above
2. What is another way to write the following equation?
 $C_2H_6(g) \rightarrow C_2H_2(g) + 2H_2(g) \Delta H = 3.2 \times 10^2 \text{ kJ/mol}$
 - a. $C_2H_6(g) \rightarrow C_2H_2(g) + 2H_2(g) + 3.2 \times 10^2 \text{ kJ}$
 - b. $C_2H_6(g) + 3.2 \times 10^2 \text{ kJ} \rightarrow C_2H_2(g) + 2H_2(g)$
 - c. $2H_2(g) + C_2H_6(g) + 3.2 \times 10^2 \text{ kJ} \rightarrow C_2H_2(g)$
 - d. $C_2H_6(g) - 3.2 \times 10^2 \text{ kJ} \rightarrow C_2H_2(g) + 2H_2(g)$
 - e. none of the above
3. 1.00 mol of an element burned in a calorimeter raise the temperature of 100 g of water by 10.0°C . The molar heat of combustion of the element is
 - a. 6.36 kJ/mol
 - b. 0.239 kJ/mol
 - c. 41.8 kJ/mol
 - d. 4.20 kJ/mol
 - e. 16.9 kJ/mol
4. Which statement concerning the accompanying diagram is true?



- a. ΔH is positive
 - b. the system is endothermic
 - c. the system releases heat to the surroundings
 - d. the heat content of the reactants is less than the heat content of the products
 - e. the enthalpy of the products is greater than the enthalpy of reactants
5. Which of the following equations represents the correct heat of formation of water vapour?
 - a. $H_2(g) + \frac{1}{2} O_2(g) \rightarrow H_2O(l)$
 - b. $H_2(g) + O(g) \rightarrow H_2O(g)$
 - c. $2H(g) + O(g) \rightarrow H_2O(g)$
 - d. $H(g) + O_2(g) \rightarrow 2H_2O(g)$
 - e. $H(g) + \frac{1}{2} O_2(g) \rightarrow H_2O(g)$

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