CHEM 1411, chapter 6. Thermochemistry Exercises

- 1. The heat capacity of 20.0 g of water is 83.7 J/°C.
 - A) True
 - B) False
- Find the heat absorbed from the surroundings when 15 g of O₂ reacts according to the equation O + O₂ → O₃. ΔH^o_{2m} = -103 kJ/mol.
 - A) 4.6 × 10 1 kJ
 - B) 48 kJ
 - C) 96 kJ
 - D) 32 kJ
 - E) 110 kJ
- 3. The heat of solution of calcium chloride CaCl₂ is -82.8 kJ/mol, and the combined heats of hydration of 1 mole of gaseous calcium ions and 2 mole of gaseous chloride ions is -2327 kJ. What is the lattice energy of calcium chloride?
- Given the following ΔH^o values.
 - $H_2(g) + \frac{1}{2}O_2(g) \rightarrow H_2O(l)$ $\Delta H^0_f = -285.8 \text{ kJ/mol}$ $H_2O_2(l) \rightarrow H_2(g) + O_2(g)$ $\Delta H^0_{loss} = 187.6 \text{ kJ/mol}$ calculate ΔH^0_{loss} for the reaction $H_2O_2(l) \rightarrow H_2O(l) + \frac{1}{2}O_2(g)$.
- 5. Calculate the amount of work done, in joules, when 2.5 mole of H₂O vaporizes at 1.0 atm and 25°C. Assume the volume of liquid H₂O is negligible compared to that of vapor. (1 L-atm = 101.3 J)
 - A) 6,190 kJ
 - B) 6.19 kJ
 - C) 61.1J
 - D) 5.66 kJ
 - E) 518 J
- The heat of solution of NH₄NO₃ is 26.2 kJ/mol. Is heat evolved or absorbed when a solution of NH₄NO₃ is diluted by addition of more water?

Thermochemistry Exercises Answers

Mr. Rohit Manglik

Thermochemistry Exercises Answers:

Ions in Solution J Burgess, 1999-10-01 This outline of the principles and chemical interactions in inorganic solution chemistry delivers a course module in an area of considerable complexity Problems with solutions and tutorial hints to test comprehension have been added as a feature to check readers understanding and assist self study Exercises and projects are also provided to help readers deepen and extend their knowledge and understanding Inorganic solution chemistry is treated thoroughly Emphasis is placed upon NMR UV VIS IR Raman spectroscopy X ray diffraction and such topics as acid base behaviour stability constants and kinetics Numerical Chemistry, *Problem Exercises for General Chemistry* G.Gilbert Long, Forrest C. Hentz, 1986-08-14 A question answer supplement to accompany a 1 year general chemistry course This text is a multiple choice format and can be used with any standard general chemistry text. The exercises emphasize the importance of problem solving and equation writing in the style used in general chemistry examinations and professional school aptitude examinations such as the MCAT and VCAT Chemistry for Nerds Guide Book: Chemistry, Science, Nerd, Geek, Textbook, Guidebook, Study Guide, Educational, STEM, Science Gift Matt Kingsley, Calling all curious minds and science enthusiasts Are you fascinated by the invisible forces that shape our world Do you crave a deeper understanding of the elements molecules and reactions that make up everything around us Then look no further than Chemistry for Nerds Unleash Your Inner Mad Scientist This isn t your typical boring textbook This is a thrilling adventure through the captivating world of chemistry written in a way that s engaging accessible and downright fun Inside these pages you ll discover The secrets of matter From atoms and molecules to the states of matter and the laws that govern them The magic of reactions Explore the explosive world of chemical reactions from kinetics and equilibrium to acids bases and buffers The wonders of the elements Unravel the mysteries of the periodic table and the trends that connect its diverse inhabitants The power of chemistry in action See how chemistry shapes our environment fuels our technologies and even sustains life itself Chemistry for Nerds is packed with Crystal clear explanations Complex concepts are broken down into bite sized pieces making even the most challenging topics easy to grasp Engaging examples and analogies Relate chemistry to everyday life with fun and memorable examples Expert practical tips Put your knowledge into action with helpful tips and tricks for mastering chemistry concepts Whether you re a student a hobbyist or simply curious about the world around you Chemistry for Nerds will ignite your passion for science and unleash your inner mad scientist Get your copy today and start exploring the amazing world of chemistry Calculations for A-level Chemistry E. N. Ramsden, 1995 Comprehensive mathematics foundation section Work on formulae and equations the mole volumetric analysis and other key areas is included Can be used as a course support book as well as for exam practice Best selling experienced chemistry author **KVPY - SA: Chemistry for** Class 11th by Career Point Kota Career Point Kota, 2020-07-26 Career Point Kota feel great pleasure to present before you this KVPY SA book Detailed Topic Wise theory supported with example Previous Year Questions Complete Solution This book

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