

EXPERIMENT 2

ELECTROCHEMISTRY: ELECTROCHEMICAL CELL AND THERMODYNAMIC FUNCTIONS

Objectives

1. To construct electrochemical cells
2. To determine the net cell potentials for three electrochemical cells
3. To calculate the net cell potential based on the potentials of the half-reactions that occur and to compare your experimental and calculated values.
4. To measure the Gibbs energy based on the cell potentials.

Introduction

The use of electrochemical cells to convert the Gibbs energy stored in the constituent half-reactions into electrical work is of enormous industrial as well as fundamental significance. We have all used batteries and these are simply galvanic cells that are similar to those you will be constructing in this experiment. In the laboratory, a typical electrochemical cell has the following general construction:

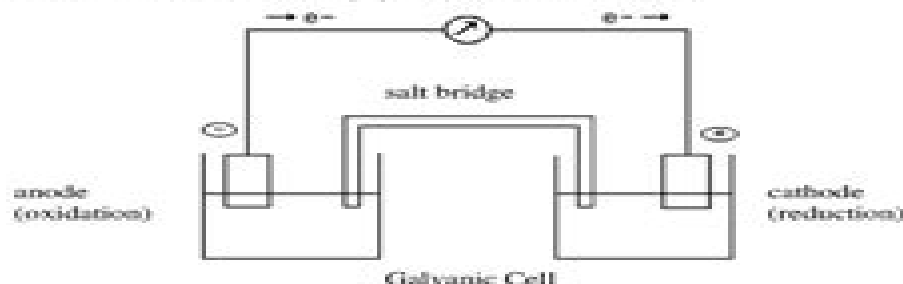


Figure 2.1: A schematic diagram of galvanic cell

The standard reduction potential is a quantitative measure of a substance's tendency to accept electrons under standard conditions (1 atm of gases and 1 M concentration) and is measured in volts. The more positive the reduction potential, the greater the tendency for reduction to occur. For the half-reaction that represents the oxidation of a metal, you can determine the standard oxidation potential (E°_{ox}) for the half-reaction. This value has the same numerical value as the standard reduction potential but the opposite sign. For example, aluminum ion has a standard reduction potential value of -1.65 V



Thermodynamics Lab Manual Experiment

Jo Allan Beran



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Laboratory Manual for Principles of General Chemistry J. A. Beran, Mark Lassiter, 2022-08-16 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition's experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section

Laboratory Manual for Principles of General Chemistry Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures

Resources in Education, 1998

Experiments in Physical Chemistry J. M. Wilson, R. J. Newcombe, A. R. Denaro, 2016-06-06 Experiments in Physical Chemistry aims to facilitate experimental work in the physical chemistry laboratory at every stage of a student's career The book is organized into three parts Part I consists of those experiments that have a simple theoretical background Part II consists of experiments that are associated with more advanced theory or more recently developed techniques or that require a greater degree of experimental skill The last part contains experiments that are in the nature of investigations This book will be useful to students to gain confidence in his ability to perform a physical chemistry experiment and to appreciate the value of the experimental approach

A Laboratory Manual of Analytical Methods of Protein Chemistry P. Alexander, H. P. Lundgren, 2014-05-17 A Laboratory Manual of Analytical Methods of Protein Chemistry Volume 5 presents the laboratory techniques for protein and polypeptide study This book discusses the staining procedure for histones which has a high degree of selectivity for basic proteins and the unique ability to visualize qualitative differences in terms of color changes Organized into four chapters this volume begins with an overview of the formalin mediated ammoniacal silver

staining procedure as a selective stain for basic proteins and its application per cell and per extract This text then examines the optical rotatory dispersion ORD which has advanced into a powerful tool for describing the conformations and conformational changes of biopolymers Other chapters consider the application of ultrasensitive calorimetry to thermodynamic problems This book discusses as well the principle of the technique its instrumentation and experimental procedures The final chapter deals with the hydrodynamic densities and preferential hydration values for protein precipitates in concentrated salt solutions This book is a valuable resource for chemists and biochemists

A Handbook of Physical Chemistry Experiments Dr. Damodar V. Prabhu, Dr. Sakina Zoher Bootwala, Prof. Anand N. Gadgil, Prof. DSc Irena Kostova, Dr. Harichandra A Parbat, Dr. Chetana M Rana, Ms. Amruta Kaskar, 2025-04-29 The importance of experimentation in Science Education cannot be overemphasised Theory can best be understood if based on experimentation Ideas and hypotheses have to be tested in the laboratory for validation and acceptance This book entitled A Handbook of Physical Chemistry Experiments discusses in detail more than 85 experiments which form a major part of the M Sc Chemistry syllabi of Indian universities and will be particularly useful to students pursuing studies in Physical Chemistry as also students pursuing equivalent studies in Chemistry in foreign universities While discussing the experiments a step by step approach has been adopted starting with the aim of the experiment theory requirements detailed procedure calculations and results Also a few experiments have been suggested without giving too many details The book deals with the experiments pertaining to Chemical kinetics Partition ratio Adsorption Electroanalytical techniques Conductometry Potentiometry pHmetry and Polarography Colorimetry and Spectrophotometry Thermochemistry Phase rule Transport numbers Cryoscopy Colloidal state of matter and Physical properties viscosity optical rotation surface tension and refractive index Appendixes on 1 Safety in the Chemistry laboratory 2 Preparation of solutions and 3 Fundamental Constants have been included A bibliography of useful reference books has been included to motivate the readers to do further reading We hope the book will be useful to students teachers researchers and industry chemists Suggestions for the improvement of the book are welcome

Physical Chemistry Laboratory Manual Ramesh Kumari, Amirtha Anand, 2018-10-05 This book covers the latest syllabus of CBCS pattern of Delhi and other universities for both B Sc Programme and Honours courses A large number of Physical Chemistry Environmental Chemistry Nanoscience Polymer Chemistry and Analytical Chemistry experiments have been covered using interdisciplinary and innovative methods The contents include some fundamental chemical concepts measurement of surface tension and viscosity colorimetry determination of order of a reaction heterogeneous equilibria adsorption on solid surfaces thermochemical measurements conductometric and potentiometric measurements pHmetry environmental parameter analysis etc Wherever possible two or more methods are given So the teachers and students will have a choice to make depending on the availability of chemicals apparatus instruments time etc This book will give them the opportunity to relate theory and practicals for a better understanding of the subject

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(U.S.),1970 **Designing and Writing the Thermodynamics Lab Manual** Carol A. Reed,1988 *The Publishers' Trade
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Table of Contents Thermodynamics Lab Manual Experiment

1. Understanding the eBook Thermodynamics Lab Manual Experiment
 - The Rise of Digital Reading Thermodynamics Lab Manual Experiment
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermodynamics Lab Manual Experiment
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermodynamics Lab Manual Experiment
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermodynamics Lab Manual Experiment
 - Personalized Recommendations
 - Thermodynamics Lab Manual Experiment User Reviews and Ratings
 - Thermodynamics Lab Manual Experiment and Bestseller Lists
5. Accessing Thermodynamics Lab Manual Experiment Free and Paid eBooks

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 - Thermodynamics Lab Manual Experiment Compatibility with Devices
 - Thermodynamics Lab Manual Experiment Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermodynamics Lab Manual Experiment
 - Highlighting and Note-Taking Thermodynamics Lab Manual Experiment
 - Interactive Elements Thermodynamics Lab Manual Experiment
- 8. Staying Engaged with Thermodynamics Lab Manual Experiment
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermodynamics Lab Manual Experiment
- 9. Balancing eBooks and Physical Books Thermodynamics Lab Manual Experiment
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermodynamics Lab Manual Experiment
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Thermodynamics Lab Manual Experiment
 - Setting Reading Goals Thermodynamics Lab Manual Experiment
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Thermodynamics Lab Manual Experiment
 - Fact-Checking eBook Content of Thermodynamics Lab Manual Experiment
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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