EXPERIMENT 8: Activity Series (Single Displacement Reactions)

PURPOSE

- a). Reactions of metals with acids and solt solutions.
- b). Determine the activity of metals
- Write a followed molecular equation, complete toxic equation, and not issue equation for single explanament reaction.

Milaterials

Equipment: 2 well-plattic (Spot plates) or enail test tubes:

Stirming road on twothspicks:

Chemicals and solutions: a tray containing the ractals Mg, Pb, Cu, Zu, Fe, 1.0 M HCl and 0.25 M aitrate solt subtriess of each of these metals such as:

Margaret Williams

Principal Charles

Charles Section 1

Zim TAKIN In

Photosophia

Note - Before you begin this lab, you should be familiar with types of reactions found in shapter 4 of the text. Wear appropriate elections, shows and sofety glasses for this lab.



Stiming rest



Specialistic:

INTRODUCTION

Our of the most important characteristics of a metal is its activity (teartivity). The activity of a metal is its ability to react with nonmentals. In such reaction a metal typically loads electrons to the new metal to form cations and the product is an ionic compound.

Different metals have different activities. Some metals are very active such as lithium, sodium, potassium, and cotions, some are slowly active such as till, lead, and tree; and some are not active at all (tractive) such as copper, silver, and recovery. Inactive metals are those which resist exhibition (reaction with oxygen in sir). These are often called the notice metals.

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LP Steffe

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