

EXPERIMENT 8.1 SINGLE-REPLACEMENT REACTIONS

Purpose:

1. To study some typical single replacement reactions.
2. To develop an activity series for a limited number of elements.

Principles:

The chemical reactivity of elements varies over a wide range. Some elements, like sodium and fluorine, are so reactive that they are never found in the free, uncombined state in nature. Other elements like carbon and platinum are nearly inert and can be made to react with other elements only under special conditions.

The reactivity of an element is related to its tendency to lose or gain electrons.

In principle, it is possible to arrange nearly all the elements into a single series in order of their reactivities. A number of tests have indicated which free elements are capable of replacing other elements from their compounds. Such a list is known as an *activity* or *electromotive* series.

To illustrate the preparation of an activity series we will experiment with a small group of selected elements and their compounds.

A generalized single replacement reaction is represented in this form:



Element A is the more active metal and replaces element B from the compound BZ. But if element B was a more active metal than element A, no reaction will occur.

Let us consider two specific examples, using copper and mercury:

Example 1

A few drops of mercury metal are added to a solution of copper (II) chloride. In this example no change is observed even after the solution has been standing for a prolonged time. Hence, we conclude that there is no reaction.



From this evidence we conclude that mercury will not replace copper in copper compounds, and therefore mercury is a less active metal than copper.

Example 2

A strip of metallic copper is immersed in a solution of mercury (II) chloride. In this example the copper strip is soon coated with metallic mercury and the solution becomes pale blue-green (the color of Cu^{2+} ions in aqueous solution).



From this evidence we conclude that copper will replace mercury in mercury compounds. Therefore copper is more reactive than mercury and is higher in the activity series.

The single replacement equation given above shows that in terms of ionic charges, the chloride ion remains unchanged, but the ionic charge of mercury has changed from $2+$ to 0 , and the ionic charge of copper has changed from 0 to $2+$.

Single Replacement Reactions Lab Answers Experiment 1

Annelies Wilder-Smith



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Table of Contents Single Replacement Reactions Lab Answers Experiment 1

1. Understanding the eBook Single Replacement Reactions Lab Answers Experiment 1
 - The Rise of Digital Reading Single Replacement Reactions Lab Answers Experiment 1
 - Advantages of eBooks Over Traditional Books
2. Identifying Single Replacement Reactions Lab Answers Experiment 1
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Single Replacement Reactions Lab Answers Experiment 1
 - User-Friendly Interface
4. Exploring eBook Recommendations from Single Replacement Reactions Lab Answers Experiment 1
 - Personalized Recommendations
 - Single Replacement Reactions Lab Answers Experiment 1 User Reviews and Ratings
 - Single Replacement Reactions Lab Answers Experiment 1 and Bestseller Lists
5. Accessing Single Replacement Reactions Lab Answers Experiment 1 Free and Paid eBooks

- Single Replacement Reactions Lab Answers Experiment 1 Public Domain eBooks
- Single Replacement Reactions Lab Answers Experiment 1 eBook Subscription Services
- Single Replacement Reactions Lab Answers Experiment 1 Budget-Friendly Options
- 6. Navigating Single Replacement Reactions Lab Answers Experiment 1 eBook Formats
 - ePub, PDF, MOBI, and More
 - Single Replacement Reactions Lab Answers Experiment 1 Compatibility with Devices
 - Single Replacement Reactions Lab Answers Experiment 1 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Single Replacement Reactions Lab Answers Experiment 1
 - Highlighting and Note-Taking Single Replacement Reactions Lab Answers Experiment 1
 - Interactive Elements Single Replacement Reactions Lab Answers Experiment 1
- 8. Staying Engaged with Single Replacement Reactions Lab Answers Experiment 1
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Single Replacement Reactions Lab Answers Experiment 1
- 9. Balancing eBooks and Physical Books Single Replacement Reactions Lab Answers Experiment 1
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Single Replacement Reactions Lab Answers Experiment 1
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Single Replacement Reactions Lab Answers Experiment 1
 - Setting Reading Goals Single Replacement Reactions Lab Answers Experiment 1
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Single Replacement Reactions Lab Answers Experiment 1
 - Fact-Checking eBook Content of Single Replacement Reactions Lab Answers Experiment 1
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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