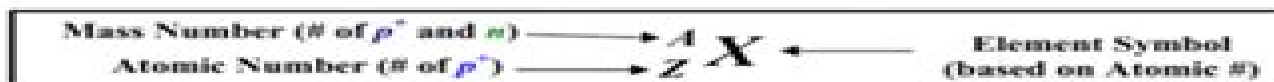


Chapter 21: Nuclear Chemistry21.1: The Nature of Nuclear Reactions

**Nucleons:** - the particles that make up a nucleus of an atom (protons, ( ${}^1_1p^+$  or  ${}^1_1H$ ) and neutrons, ( ${}^1_0n$ )).

**Isotopes:** - atoms that have different mass number but the same atomic number or number of protons.

**Nuclide:** a particular atom or isotope containing specific numbers of protons and neutrons



**Radioactivity:** - the particles and/or electromagnetic radiation that are emitted due to unstable nuclei.  
 - all elements having atomic number 84 (Polonium) and greater are radioactive.

**Nuclear Transmutation:** - a process where radioactivity is resulted from the bombardment of nuclei by neutrons, protons or other nuclei.  
 - in most cases, heavier elements are synthesized from lighter elements.

History of Radioactivity

- In 1896, **Wilhelm Roentgen** discovered X-ray by examining ray emitting from the outside of the cathode ray glass tube. It has the capability of passing through solid materials, but can be blocked by denser matters. It can also be exposed to photographic plate, resulting from an image of "seeing through" an container made of less dense material.
- In the same year, **Antoine Becquerel** discovered that uranium emits a ray onto a photographic plate in the absence of sunlight or other forms of energy.
- A few years later, **Marie and Pierre Curie** demonstrated that radiation can be emitted by other elements. They discovered two new elements (polonium and radium) based on their tendency to emit radiation (**radioactivity**).

Types of Radioactive Particle and Decay

- Alpha Particle ( $\alpha$  particle):** - basically a helium nucleus ( ${}^4_2He$ ), commonly found during radioactive decay from heavier nuclide (the net result is to **increase the neutron to proton ratio** – more explanation in the next section).



- Beta Particle ( $\beta$  particle):** - basically an electron ( ${}^0_{-1}e$  or  ${}^0_{-1}\beta$ ) that is emitted when the neutron to proton ratio is higher than the zone of stability (**a neutron is transformed to a proton and an electron as a result** – more explanation in the next section).  
 - electrons have a mass number of 0 and an atomic number assignment of -1, due to its charge.



# Topic 12 Nuclear Chemistry Review Questions Answers

**Richard Bailey**



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### **Table of Contents Topic 12 Nuclear Chemistry Review Questions Answers**

1. Understanding the eBook Topic 12 Nuclear Chemistry Review Questions Answers
  - The Rise of Digital Reading Topic 12 Nuclear Chemistry Review Questions Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Topic 12 Nuclear Chemistry Review Questions Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Topic 12 Nuclear Chemistry Review Questions Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Topic 12 Nuclear Chemistry Review Questions Answers
  - Personalized Recommendations
  - Topic 12 Nuclear Chemistry Review Questions Answers User Reviews and Ratings
  - Topic 12 Nuclear Chemistry Review Questions Answers and Bestseller Lists

5. Accessing Topic 12 Nuclear Chemistry Review Questions Answers Free and Paid eBooks
  - Topic 12 Nuclear Chemistry Review Questions Answers Public Domain eBooks
  - Topic 12 Nuclear Chemistry Review Questions Answers eBook Subscription Services
  - Topic 12 Nuclear Chemistry Review Questions Answers Budget-Friendly Options
6. Navigating Topic 12 Nuclear Chemistry Review Questions Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Topic 12 Nuclear Chemistry Review Questions Answers Compatibility with Devices
  - Topic 12 Nuclear Chemistry Review Questions Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Topic 12 Nuclear Chemistry Review Questions Answers
  - Highlighting and Note-Taking Topic 12 Nuclear Chemistry Review Questions Answers
  - Interactive Elements Topic 12 Nuclear Chemistry Review Questions Answers
8. Staying Engaged with Topic 12 Nuclear Chemistry Review Questions Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Topic 12 Nuclear Chemistry Review Questions Answers
9. Balancing eBooks and Physical Books Topic 12 Nuclear Chemistry Review Questions Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Topic 12 Nuclear Chemistry Review Questions Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Topic 12 Nuclear Chemistry Review Questions Answers
  - Setting Reading Goals Topic 12 Nuclear Chemistry Review Questions Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Topic 12 Nuclear Chemistry Review Questions Answers
  - Fact-Checking eBook Content of Topic 12 Nuclear Chemistry Review Questions Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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