## Periodic Trends

Can the properties of an element be predicted using a periodic table?

# Why?

The periodic table is often considered to be the "best friend" of chemists and chemistry students alike. It includes information about atomic masses and element symbols, but it can also be used to make predictions about atomic size, electronegativity, ionization energies, bonding, solubility, and reactivity. In this activity you will look at a few periodic trends that can help you make those predictions. Like most trends, they are not perfect, but useful just the same.

- Consider the data in Model 1 on the following page.
  - a. Each element has three numbers listed under it. Which value represents the atomic radius?

    The first number below the diagram
  - b. What are the units for the atomic radius?

pm picometers 10-11 m

e. Write a complete sentence to convey your understanding of atomic radius. Nov. You many not use the word "radius" in your definition.

Distance from center of atom (nucleus) to outer elect rons

 In general, what is the trend in atomic radius as you go down a group in Model 1? Support your answer, using examples from three groups.

Incremes H<Li< Na < K

FKCIKBA

Using your knowledge of Coulombic attraction and the structure of the atom, explain the trend
in atomic radius that you identified in Question 2. Hine: You should discuss either a change in
distance between the nucleus and outer shell of electrons or a change in the number of protons
in the nucleus.

As the number of electrons sucremes, now energy levels need to be added first her from the nucleus, increasing the radius

 In general, what is the trend in atomic radius as you go across a period (left to right) in Model 1? Support your answer, using examples from two periods.

Decreases Nasmy > Alssisps son > Ar

 Using your knowledge of Coulombic attraction and the structure of the atom, explain the trend in atomic radius that you identified in Question 4.

the ancient increases, increasing the charge of the ancient and the attractive force on the electrons, pulling the highest Periodic Trends energy level closer to the nucleus

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**SA Dillow** 

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#### **Table of Contents Trands Packet Answers**

- 1. Understanding the eBook Trands Packet Answers
  - The Rise of Digital Reading Trands Packet Answers
  - o Advantages of eBooks Over Traditional Books
- 2. Identifying Trands Packet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Trands Packet Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Trands Packet Answers
  - Personalized Recommendations
  - Trands Packet Answers User Reviews and Ratings
  - Trands Packet Answers and Bestseller Lists
- 5. Accessing Trands Packet Answers Free and Paid eBooks
  - Trands Packet Answers Public Domain eBooks
  - Trands Packet Answers eBook Subscription Services

- Trands Packet Answers Budget-Friendly Options
- 6. Navigating Trands Packet Answers eBook Formats
  - ∘ ePub, PDF, MOBI, and More
  - Trands Packet Answers Compatibility with Devices
  - Trands Packet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Trands Packet Answers
  - Highlighting and Note-Taking Trands Packet Answers
  - Interactive Elements Trands Packet Answers
- 8. Staying Engaged with Trands Packet Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Trands Packet Answers
- 9. Balancing eBooks and Physical Books Trands Packet Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Trands Packet Answers
- 10. Overcoming Reading Challenges
  - $\circ$  Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Trands Packet Answers
  - Setting Reading Goals Trands Packet Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Trands Packet Answers
  - Fact-Checking eBook Content of Trands Packet Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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