

Chemistry Honors Final Study Guide

- Semester 2 -

Solids, Liquids, and Solutions | Chapters 14 & 15:

⇒ Intermolecular Forces

- **Intermolecular Forces:** forces that hold molecules together in liquids and solids (are weaker than intramolecular forces)
 - Types of intramolecular forces are ionic bonds, covalent bonds, and metallic bonds.
 - The 3 intermolecular forces are *dipole-dipole forces*, *dispersion forces*, and *hydrogen bonding*.

✧ Dipole-Dipole Forces

- **Dipole-Dipole Forces:** attractions between opposite charges of neighboring (permanent) dipoles.
- Polar molecules are attracted to opposite ends of each other through dipole-dipole forces
- This intermolecular force is only possible with polar molecules

✧ Ion-Dipole Forces

- **Ion Dipole Force:** an attractive force that results from the electrostatic attraction between an ion and a neutral molecule that has a dipole

✧ Dispersion Forces

- **Induced Dipole:** a dipole created by the presence of a neighboring dipole
- **Dispersion Forces:** the force of attraction between induced dipoles (occur whenever electrons are involved)
- **Polarizability:** the ease with which the electron cloud of an atom or molecule is distorted
 - The respective trend is atomic (or molar) mass/radius/weight and it is the exact opposite of the ΔEN trend on the periodic table

Semester 2 Chemistry Honors Review Guide

JE Gale



Semester 2 Chemistry Honors Review Guide:

Semester 2 Chemistry Honors Review Guide Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Semester 2 Chemistry Honors Review Guide**," written by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve to the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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