

## Practical 2 - Qualitative Analysis of Cations & Anions I

Sample 1 (cation group I)	
Sample name and formula	Silver Nitrate ( )
Precipitation test reagent	1. Add 2ml of Dilute Hydrochloric Acid (HCl) into a test tube and label 2. Add Ammonium Hydroxide (OH)
Precipitate identification	1. Precipitate is a white solid 2. Precipitate dissolves (is a colorless solution) ( )
1 <sup>st</sup> confirmatory test	Add Potassium Iodide (KI) to the solution. Turns Light yellowish in color
2 <sup>nd</sup> confirmatory test	N/A
Anion test and identification	<p>Chloride Anion</p> <ul style="list-style-type: none"> <li>Add 2ml of Dilute Hydrochloric acid into a clean test tube.</li> <li>Carefully Acidify the solution by adding a few drops of Nitric Acid ( )</li> <li>Add a few drops of Silver Nitrate ( )</li> <li>The formation of a white precipitate (AgCl) that dissolves in dilute OH (to form a colorless indicates the presence of the chloride anion ).</li> </ul>

Sample 2 (cation group II)	
Sample name and formula	Copper Sulfate
Precipitation test reagent	<ol style="list-style-type: none"> <li>Add 2ml of the test solution to a clean test tube.</li> <li>Carefully add 2ml of dilute hydrochloric acid( )</li> <li>In the fume hood add Thioacetamide ( ) dropwise this is used as a source of Hydrogen Sulfide ( )</li> <li>Heat gently in the steam bath in the fume hood to allow the Thioacetamide to hydrolyze slowly to produce ( )</li> </ol>
Precipitate identification	Black precipitate formed due to the presence of ( )
1 <sup>st</sup> confirmatory test	Sodium Hydroxide Solution ( ) turns black/ pale blue in color.
2 <sup>nd</sup> confirmatory test	Potassium Ferrocyanide solution ( ) turns a dark redish brown color
Anion test and identification	<p>Nitrate Anion</p> <ol style="list-style-type: none"> <li>Add 2ml of the test solution to a clean test tube.</li> <li>Acidify by adding some drops of dilute ( ).</li> </ol>

# Separation And Analysis Of Cations Practical Report

**M Woodhall**



## **Separation And Analysis Of Cations Practical Report:**

ASD Technical Report United States. Air Force. Systems Command. Aeronautical Systems Division, 1961      SARAP Research Technical Reports ,1984      *Accessions of Unlimited Distribution Reports* ,1974-07-05      *Scientific and Technical Aerospace Reports* ,1994      **Nuclear Science Abstracts** ,1974      **Energy Research Abstracts** ,1985      Selected Water Resources Abstracts ,1988      Handbook of Capillary Electrophoresis James P. Landers,1994 The Handbook of Capillary Electrophoresis is a valuable collection of in depth discussions of specific areas of capillary electrophoresis CE that are key to the practical use of CE for sample analysis The handbook is divided into five parts Part I focuses on various modes of CE that have been developed and applied to analytical separation problems Part II covers detection in CE with one section reviewing all detector systems that have been successfully interfaced with CE and another that focuses on CE mass spectrometry Part III explores how CE can be applied to the analysis of a specific class of molecules analytes or to solving separation problems CE analysis of organic ions pharmaceuticals glycoconjugates peptides proteins and DNA fragments the components of single cells and the utility of capillary coatings in CE are all covered in this section A key element of this handbook is its discussion of the potential role of CE for solving clinical problems and the potential for CE and CE MS to advance analysis of drug metabolism The handbook also presents practical and theoretical discussions on various aspects of CE including sample matrix and its effects on separation the need for controlling capillary temperature and the present and future implications of EOF control In depth discussions of this nature comprise Parts IV and V The Handbook of Capillary Electrophoresis will be an outstanding reference for analytical chemists biochemists clinical chemists cell biologists and pharmaceutical researchers

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