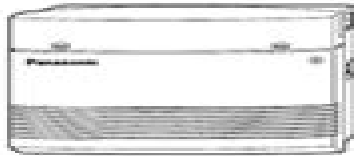


# Panasonic



## Advanced Hybrid System User Manual

Model No. **KX-TA624**



Please read this manual before using the Advanced Hybrid System.

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**Ying-Ying Zheng**



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in the n 2 molecule the nitrogen atoms have an  $\sigma$  bond and two  $\pi$  bonds holding the two atoms together the presence of three

strong bonds makes  $n = 2$  a very stable molecule phosphorus is a third period element and as such does not form  $\pi$  bonds efficiently therefore it must fulfill its bonding requirement by forming three  $\sigma$  bonds

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complete atoms can or electrons to get a stable outer energy level gain lose share a figure that uses the symbol of the element and dots to represent the electrons in the outer energy level also used to show how valence electrons are bonded when elements combine to form compounds

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