

## [Section 10.1]

# Ions in Aqueous Solutions and Colligative Properties

## Section Quiz: Colligative Properties of Solutions

Write each problem within the outline of the quiz or phrase that best completes each statement or best answers each question.

- \_\_\_\_\_ 1. Which of the following *is not* considered a colligative property?
- It is directly proportional to the number of particles in solution.
  - It is directly proportional to the number of particles in solution.
  - It depends on the properties of an individual solute molecule.
  - None of the above.
- \_\_\_\_\_ 2. Compounds with a total of 4  $\text{Cl}^-$ , 2  $\text{Ca}^{2+}$ , and 10  $\text{Na}^+$  ions in a solution give \_\_\_\_\_.
- the same freezing-point depression.
  - about twice the freezing-point depression.
  - the same freezing-point elevation.
  - about one twice the freezing-point elevation.
- \_\_\_\_\_ 3. Compound with a total of four particles, a total of 10  $\text{Na}^+$ , 2  $\text{Ca}^{2+}$ , and 10  $\text{Cl}^-$  ions \_\_\_\_\_.
- the same freezing-point depression.
  - about half of the freezing-point elevation.
  - is freezing-point elevation about two-thirds of 10.
  - is freezing-point elevation about 1.5 times of 10.
- \_\_\_\_\_ 4. When a nonelectrolyte solute dissolved in water, the liquid portion of the solution \_\_\_\_\_.
- increases.
  - decreases.
  - does not change.
  - changes depending on the solute used.
- \_\_\_\_\_ 5. Which of the following compounds would be most effective in lowering the boiling point of an aqueous solution?
- $\text{C}_2\text{H}_6\text{O}$
  - $\text{H}_2\text{SO}_4$
  - $\text{H}_2\text{O}$
  - $\text{H}_2\text{SO}_3$

# Rosengarten Colligative Properties Homework Answers

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