

Key

Ideal Gas Law Worksheet $PV = nRT$

Use the ideal gas law, "PerV-nRT", and the universal gas constant $R = 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}}$ to solve the following problems:

If pressure is needed in kPa then convert by multiplying by $101.3 \text{ kPa} / 1 \text{ atm}$ to get
 $R = 8.31 \text{ kPa} \cdot \text{L} / (\text{K} \cdot \text{mole})$

- 1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?

$$PV = nRT$$

$$T = \frac{PV}{nR} = \frac{(5.6 \text{ atm})(12 \text{ L})}{4 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}}}$$

$$T = 204.63 \text{ K}$$

- 2) If I have an unknown quantity of gas at a pressure of 1.2 atm, a volume of 31 liters, and a temperature of 87°C , how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{1.2 \text{ atm} \cdot 31 \text{ L}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 360 \text{ K}}$$

$$n = 1.2586 \text{ mol}$$

- 3) If I contain 3 moles of gas in a container with a volume of 60 liters and at a temperature of 400 K, what is the pressure inside the container?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{3 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 400 \text{ K}}{60 \text{ L}}$$

$$P = 1.642 \text{ atm}$$

or $P = 166.29 \text{ kPa}$

- 4) If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of 56°C ,

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{7.7 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 329 \text{ K}}{0.09 \text{ atm}}$$

$$V = 2916.1 \text{ L}$$

- 5) If I have an unknown quantity of gas at a pressure of 0.75 atm, a volume of 100 liters, and a temperature of 30°C , how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{0.75 \text{ atm} \cdot 100 \text{ L}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 303 \text{ K}}$$

$$n = 2.91 \text{ mol}$$

$$n = 2.91 \text{ mol}$$

- 6) If I have 2.5 moles of gas at a pressure of 0.85 atm, a volume of 100 liters, and a temperature of 30°C , what is the pressure inside the container?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{2.5 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 303 \text{ K}}{100 \text{ L}}$$

$$P = 0.62 \text{ atm}$$

Weston Walch Gas Law Problems

Didier Musso



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